

U.S. DEPARTMENT OF COMMERCE PATENT & TRADEMARK OFFICE

B/O Form PTO-1390	Transmittal Letter to the United States Designated/Elected Office (DO/EO/US) Concerning a Filing Under 35 USC 371	Attorney's Docket Number PICK3002/REF/571
International Application Number PCT/GB00/00571		U.S. Application Number (if known) 09/926051
Title of Invention STABILIZATION OF TRANSITION ALUMINA	International Filing Date 17 February 2000	Priority Date Claimed 24 February 1999
Applicant(s) for DO/EO/US PICKERING et al.		

Applicant herewith submits to the United States Designated/Elected Office (DO/EO/US) the following items under 35 USC 371:

1. ☒ This is a **FIRST** submission of items concerning a filing under 35 USC 371.
2. ☐ This is a **SECOND** or **SUBSEQUENT** submission of items concerning a filing under 35 USC 371.
3. ☒ This express request to begin national examination procedures (35 USC 371(f)) at any time rather than delay examination until the expiration of the applicable time limit set in 35 USC 371(b) and PCT Articles 22 and 39(1).
4. ☒ A proper Demand for International Preliminary Examination was made by the 19th month from the earliest claimed priority date.
5. ☒ A copy of the International Application as filed 35 USC 371(c)(2).
 - a. ☒ is transmitted herewith (required only if not transmitted by the International Bureau).
 - b. ☐ has been transmitted by the International Bureau.
 - c. ☐ is not required, as the application was filed in the United States Receiving Office (RO/US).
6. ☐ A translation of the International Application into English (35 USC 371(c)(2)).
7. ☒ Amendments to the claims of the International Application under PCT Article 19 (35 USC 371(c)(3))
 - a. ☐ are transmitted herewith (required only if not transmitted by the International Bureau).
 - b. ☐ have been transmitted by the International Bureau.
 - c. ☐ have not been made; however, the time limit for making such amendments has NOT expired.
 - d. ☒ have not been made and will not be made.
8. ☐ A translation of the amendments to the claims under PCT Article 19 (35 USC 371(c)(3)).
9. ☐ An oath or declaration of the inventor(s) (35 USC 371(c)(4)). (☐ Executed ☐ Unexecuted)
10. ☐ A translation of the annexes to the International Preliminary Examination Report under PCT Article 36 (35 USC 371(c)(5)).

Items 11 to 16 below concern other document(s) or information included:

11. ☒ An Information Disclosure Statement under 37 CFR 1.97 and 1.98.
12. ☐ An assignment document for recording. A separate cover sheet in compliance with 37 CFR 3.28 and 3.31 is included.
13. ☒ A **FIRST** preliminary amendment.
☐ A **SECOND** or **SUBSEQUENT** preliminary amendment.
14. ☐ A substitute specification.
15. ☐ A change of power of attorney and/or address letter.
16. ☒ Other items or information: Applicants assert entitlement to small entity status.

Application Number (if Known) 09/926051		International Application Number PCT/GB00/00571		Attorney's Docket Number PICK3002/REF/571	
				Calculations	PTO USE ONLY
17. The following fees are submitted: Basic National Fee (37 CFR 1.492(a)(1)-(5)): <input checked="" type="checkbox"/> Search report has been prepared by the EPO or JPO \$860.00 <input type="checkbox"/> International Preliminary Examination Fee paid to USPTO (37 CFR 1.482) \$690.00 <input type="checkbox"/> No International Preliminary Examination Fee paid to USPTO (37 CFR 1.482) but International Search Fee paid to USPTO (37 CFR 1.445(a)(2)) \$710.00 <input type="checkbox"/> Neither International Preliminary Examination Fee (37 CFR 1.482) nor International Search Fee (37 CFR 1.445(a)(2)) paid to USPTO \$1000.00 <input type="checkbox"/> International Preliminary Examination Fee paid to USPTO (37 CFR 1.482) and all claims satisfied provisions of PCT Article 33(1)-(4) \$100.00				\$860.00	
ENTER APPROPRIATE BASIC FEE AMOUNT				\$ 860.00	
Surcharge of \$130.00 for furnishing the oath or declaration later than <input type="checkbox"/> 20 <input type="checkbox"/> 30 months from the earliest claimed priority date (37 CFR 1.492(e)).					
CLAIMS	NUMBER FILED	NUMBER EXTRA	RATE		
Total Claims	27 -20 =	7	× \$18.00	\$ 126.00	
Independent Claims	3 -3 =	0	× \$80.00	\$ 0.00	
Multiple Dependent Claims (if applicable)			+ \$270.00		
TOTAL OF ABOVE CALCULATIONS				\$ 986.00	
Reduction by ½ for filing by small entity, if applicable. Small Entity Status is asserted pursuant to 37 CFR 1.27 for this application.				\$ 493.00	
SUBTOTAL				\$ 493.00	
Processing fee of \$130.00 for furnishing the English translation later than <input type="checkbox"/> 20 <input type="checkbox"/> 30 months from the earliest claimed priority date (37 CFR 1.492(f)).					
TOTAL NATIONAL FEE					
Fee for recording the enclosed assignment (37 CFR 1.21(h)). The assignment must be accompanied by an appropriate cover sheet (37 CFR 3.28, 3.31). \$40.00 per property.					
TOTAL FEES ENCLOSED				\$ 493.00	
				Amount to be:	Refunded:
					Charged:

- a. ☒ A check in the amount of **\$493.00** to cover the fees is enclosed.
- b. ☐ Please charge my **Deposit Account Number 02-0200** in the amount of \$_____ to cover the above fees.
A duplicate copy of this sheet is enclosed.
- c. ☒ The Commissioner is hereby authorized to charge any additional fees which may be required, or credit any overpayment to **Deposit Account Number 02-0200**. A duplicate copy of this sheet is enclosed.

Note: Where an appropriate time limit under 37 CFR 1.494 or 1.495 has not been met, a petition to revive (37 CFR 1.137(a) or (b)) must be filed and granted to restore the application to pending status.

BACON & THOMAS, PLLC
 625 SLATERS LANE - FOURTH FLOOR
 ALEXANDRIA, VIRGINIA 223124-1176
 (703) 683-0500

Respectfully submitted,

Richard E. Fichter

Richard E. Fichter
 Attorney for Applicant
 Registration Number: 26,382

DATE: August 22, 2001

U.S. DEPARTMENT OF COMMERCE PATENT & TRADEMARK OFFICE

B/O Form PTO-1390		Transmittal Letter to the United States Designated/Elected Office (DO/EO/US) Concerning a Filing Under 35 USC 371		Attorney's Docket Number PICK3002/REF/571	
				U.S. Application Number (if known) 09/926051	
International Application Number PCT/GB00/00571		International Filing Date 17 February 2000		Priority Date Claimed 24 February 1999	
Title of Invention STABILIZATION OF TRANSITION ALUMINA					
Applicant(s) for DO/EO/US PICKERING et al.					

Applicant herewith submits to the United States Designated/Elected Office (DO/EO/US) the following items under 35 USC 371:

1. ☒ This is a **FIRST** submission of items concerning a filing under 35 USC 371.
2. ☐ This is a **SECOND** or **SUBSEQUENT** submission of items concerning a filing under 35 USC 371.
3. ☒ This express request to begin national examination procedures (35 USC 371(f)) at any time rather than delay examination until the expiration of the applicable time limit set in 35 USC 371(b) and PCT Articles 22 and 39(1).
4. ☒ A proper Demand for International Preliminary Examination was made by the 19th month from the earliest claimed priority date.
5. ☒ A copy of the International Application as filed 35 USC 371(c)(2).
 - a. ☒ is transmitted herewith (required only if not transmitted by the International Bureau).
 - b. ☐ has been transmitted by the International Bureau.
 - c. ☐ is not required, as the application was filed in the United States Receiving Office (RO/US).
6. ☐ A translation of the International Application into English (35 USC 371(c)(2)).
7. ☒ Amendments to the claims of the International Application under PCT Article 19 (35 USC 371(c)(3))
 - a. ☐ are transmitted herewith (required only if not transmitted by the International Bureau).
 - b. ☐ have been transmitted by the International Bureau.
 - c. ☐ have not been made; however, the time limit for making such amendments has NOT expired.
 - d. ☒ have not been made and will not be made.
8. ☐ A translation of the amendments to the claims under PCT Article 19 (35 USC 371(c)(3)).
9. ☐ An oath or declaration of the inventor(s) (35 USC 371(c)(4)). (☐ Executed ☐ Unexecuted)
10. ☐ A translation of the annexes to the International Preliminary Examination Report under PCT Article 36 (35 USC 371(c)(5)).

Items 11 to 16 below concern other document(s) or information included:

11. ☒ An Information Disclosure Statement under 37 CFR 1.97 and 1.98.
12. ☐ An assignment document for recording. A separate cover sheet in compliance with 37 CFR 3.28 and 3.31 is included.
13. ☒ A **FIRST** preliminary amendment.
 - ☐ A **SECOND** or **SUBSEQUENT** preliminary amendment.
14. ☐ A substitute specification.
15. ☐ A change of power of attorney and/or address letter.
16. ☒ Other items or information: Applicants assert entitlement to small entity status.

Application Number (if Known) 09/926051		International Application Number PCT/GB00/00571		Attorney's Docket Number PICK3002/REF/571	
				Calculations	PTO USE ONLY
17. The following fees are submitted: Basic National Fee (37 CFR 1.492(a)(1)-(5)): <input checked="" type="checkbox"/> Search report has been prepared by the EPO or JPO \$860.00 <input type="checkbox"/> International Preliminary Examination Fee paid to USPTO (37 CFR 1.482) \$690.00 <input type="checkbox"/> No International Preliminary Examination Fee paid to USPTO (37 CFR 1.482) but International Search Fee paid to USPTO (37 CFR 1.445(a)(2)) \$710.00 <input type="checkbox"/> Neither International Preliminary Examination Fee (37 CFR 1.482) nor International Search Fee (37 CFR 1.445(a)(2)) paid to USPTO \$1000.00 <input type="checkbox"/> International Preliminary Examination Fee paid to USPTO (37 CFR 1.482) and all claims satisfied provisions of PCT Article 33(1)-(4) \$100.00				\$860.00	
ENTER APPROPRIATE BASIC FEE AMOUNT				\$ 860.00	
Surcharge of \$130.00 for furnishing the oath or declaration later than <input type="checkbox"/> 20 <input type="checkbox"/> 30 months from the earliest claimed priority date (37 CFR 1.492(e)).					
CLAIMS	NUMBER FILED	NUMBER EXTRA	RATE		
Total Claims	27 -20 =	7	× \$18.00	\$ 126.00	
Independent Claims	3 -3 =	0	× \$80.00	\$ 0.00	
Multiple Dependent Claims (if applicable)			+ \$270.00		
TOTAL OF ABOVE CALCULATIONS				\$ 986.00	
Reduction by ½ for filing by small entity, if applicable. Small Entity Status is asserted pursuant to 37 CFR 1.27 for this application.				\$ 493.00	
SUBTOTAL				\$ 493.00	
Processing fee of \$130.00 for furnishing the English translation later than <input type="checkbox"/> 20 <input type="checkbox"/> 30 months from the earliest claimed priority date (37 CFR 1.492(f)).					
TOTAL NATIONAL FEE					
Fee for recording the enclosed assignment (37 CFR 1.21(h)). The assignment must be accompanied by an appropriate cover sheet (37 CFR 3.28, 3.31). \$40.00 per property.					
TOTAL FEES ENCLOSED				\$ 493.00	
				Amount to be:	Refunded:
					Charged:

- a. ☒ A check in the amount of **\$493.00** to cover the fees is enclosed.
- b. ☐ Please charge my **Deposit Account Number 02-0200** in the amount of \$_____ to cover the above fees.
A duplicate copy of this sheet is enclosed.
- c. ☒ The Commissioner is hereby authorized to charge any additional fees which may be required, or credit any overpayment to **Deposit Account Number 02-0200**. A duplicate copy of this sheet is enclosed.

Note: Where an appropriate time limit under 37 CFR 1.494 or 1.495 has not been met, a petition to revive (37 CFR 1.137(a) or (b)) must be filed and granted to restore the application to pending status.

BACON & THOMAS, PLLC
 625 SLATERS LANE - FOURTH FLOOR
 ALEXANDRIA, VIRGINIA 223124-1176
 (703) 683-0500

DATE: August 22, 2001

Respectfully submitted,

Richard E. Fichter

Richard E. Fichter
 Attorney for Applicant
 Registration Number: 26,382

PATENT

IN THE UNITED STATES PATENT AND TRADEMARK OFFICE

In re application of:

PICKERING et al.

U.S. National Phase of PCT/GB00/00571

Entry papers filed herewith August 22, 2001

For: STABILIZATION OF TRANSITION ALUMINA

Attention: PCT OFFICE

**PRELIMINARY AMENDMENT
AND INFORMATION DISCLOSURE STATEMENT**

Assistant Commissioner for Patents
Washington, D.C. 20231

Sir:

The present application is the U.S. national phase of international application number PCT/GB00/00571. The following amendments pertain to the claims as amended.

Please note that the amended pages 30-34 attached to the International Preliminary Examination Report (Annexes) and submitted herewith, have replaced the originally filed pages 30-34 of the application. The claims to be examined and amended by this preliminary amendment are found on amended pages 30-34.

Please amend the above-identified application as follows:

IN THE SPECIFICATION:

Please add the attached ABSTRACT OF THE DISCLOSURE to the application.

IN THE CLAIMS:

Please replace claims 3-5, 9-11, 13-16, 18, 20-22 and 27 with the following amended claims.

3(Amended). A process as claimed in claim 1, wherein the aluminium containing compound comprises aluminium chloride, aluminium nitrate, aluminium sulphate, alumina and/or boehmite.

4(Amended). A process as claimed in claim 1, wherein the barium containing compound comprises barium chloride, barium sulphate and/or barium nitrate.

5(Amended). A process as claimed in claim 1, wherein the mixture comprising aluminium and barium containing compounds is in the form of a substantially homogeneous precipitate.

9(Amended). A process as claimed in claim 6, wherein the salt solution further includes a dispersant, such as polyvinylpyrrolidone.

10(Amended). A process as claimed in claim 1, wherein the mixture in step (a) is formed by adding a solution comprising oxalic acid and a water soluble barium compound, preferably barium nitrate, to an aqueous slurry comprising boehmite.

11(Amended). A process as claimed in claim 1, wherein heating of the mixture in step (b) is achieved by a hydrothermal process.

13(Amended). A process as claimed in claim 11, wherein the hydrothermal treatment is carried out at a pressure in the range of from 1 to 150 bar, preferably from 5 to 50 bar.

14(Amended). A process as claimed in claim 11, wherein the hydrothermal treatment is carried out for a time in the range of from 30 minutes to 25 hours, preferably from 1 to 10 hours.

15(Amended). A process as claimed in claim 1, wherein the mixture in step (b) is heated at a temperature in the range of from ambient temperature to boiling temperature, preferably by a reflux process.

16(Amended). A process for the preparation of a stabilized alumina catalyst support material, which process comprises the steps of:

(i) providing a precursor material comprising particles of boehmite at least partially coated with a layer comprising a compound of barium, wherein the precursor material is produced according to a process as claimed in claim 1; and

(ii) heating the coated particles of boehmite to a temperature at which least some of the boehmite transforms to gamma-alumina.

18(Amended). A process for the preparation of an automotive catalytic converter, which process comprises the steps of:

(i) providing a precursor material comprising particles of boehmite at least partially coated with a layer comprising a compound of barium, wherein the precursor material is produced according to a process as claimed in claim 1;

(ii) dispersing a mixture comprising said precursor material on a metal or ceramic monolithic substrate; and

(iii) heating said mixture comprising said precursor material to a temperature at which at least some of the boehmite transforms to gamma-alumina.

20(Amended). A process as claimed in claim 16, wherein heating is carried out at a temperature in the range of from 500 to 1000°C.

21(Amended). A process as claimed in claim 17, wherein the further heating is carried out at a temperature in the range of from 1000 to 1400°C.

22(Amended). A process for the preparation of a catalyst, such as an automotive catalytic converter, which process comprises the step of dispersing a stabilized alumina

catalyst support material prepared according to the process of claim 16 on a metal or ceramic monolithic substrate.

27(Amended). A stabilized alumina catalyst carrier or support material as claimed in claim 25 having a specific surface area of at least $31 \text{ m}^2/\text{g}$ after exposure to a temperature of 1400°C 1 hour.

REMARKS

Applicants have amended the claims in order to reduce the initial filing fee by deleting the multiple dependent claims from the application. Applicants retain the right to reintroduce any subject matter canceled by the present Amendment at any time during the prosecution of this application or any further application claiming benefit of this application.

Applicants have amended the application to substitute the originally filed pages 30 - 34 with the amended pages 30 - 34 attached to the International Preliminary Examiner Report (Annexes) and included in the application as filed herewith. Also, an Abstract of the Disclosure has been added to the application.


Applicants are submitting herewith a copy of the Search Report which issued on International Application No. pct/gb00/00571, of which the present application is the U.S. national phase. All of the publications cited in the International Search Report are listed on the attached Form PTO-1449. It is Applicants' understanding that, under the procedures of the PCT, copies of the cited publications will have been supplied to the U.S. Patent Office by the International Bureau. However, the Examiner is invited to contact the undersigned attorney if additional copies are necessary or would facilitate examination of the present application.

Otherwise, the Examiner is respectfully requested to return an initialed and dated copy of the attached Form PTO-1449 to confirm that all publications listed thereon have been considered and made officially of record in the file of this application.

Applicants understand that, under the procedures of the PCT, a copy of the priority document (GB 9904269.9, filed 24 February 1999) will have been supplied to the U.S. Patent Office pursuant to Rule 17 of the PCT Regulations. It is therefore respectfully requested that the first Official Action in the present application contain an indication that the appropriate priority document is in the file of this application.

In view of the above amendments, an early action on the application is now in order and is most respectfully requested.

Respectfully submitted,
BACON & THOMAS, PLLC

By: 
RICHARD E. FICHTER
Registration No. 26,382

625 Slaters Lane - 4th Floor
Alexandria, Virginia 22314
Phone: (703) 683-0500
Facsimile: (703) 683-1080

REF:kdd
PA01.wpd

DATE: August 22, 2001

0926051.11501
T05111.15092660

Marked-Up Version Showing Changes Made

IN THE CLAIMS:

Please replace claims 3-5, 9-11, 13-16, 18, 20-22 and 27 with the following amended claims.

3(Amended). A process as claimed in claim 1 [or claim 2], wherein the aluminium containing compound comprises aluminium chloride, aluminium nitrate, aluminium sulphate, alumina and/or boehmite.

4(Amended). A process as claimed in [any one of the preceding claims] claim 1, wherein the barium containing compound comprises barium chloride, barium sulphate and/or barium nitrate.

5(Amended). A process as claimed in [any one of the preceding claims] claim 1, wherein the mixture comprising aluminium and barium containing compounds is in the form of a substantially homogeneous precipitate.

9(Amended). A process as claimed in [any one of claims 6 to 8] claim 6, wherein the salt solution further includes a dispersant, such as polyvinylpyrrolidone.

10(Amended). A process as claimed in [any one of claims 1 to 4] claim 1, wherein the mixture in step (a) is formed by adding a solution comprising oxalic acid and a water soluble barium compound, preferably barium nitrate, to an aqueous slurry comprising boehmite.

11(Amended). A process as claimed in [any one of the preceding claims] claim 1, wherein heating of the mixture in step (b) is achieved by a hydrothermal process.

13(Amended). A process as claimed in claim 11 [or claim 12], wherein the hydrothermal treatment is carried out at a pressure in the range of from 1 to 150 bar, preferably from 5 to 50 bar.

14(Amended). A process as claimed in [any one of claims 11 to 13] claim 11, wherein the hydrothermal treatment is carried out for a time in the range of from 30 minutes to 25 hours, preferably from 1 to 10 hours.

15(Amended). A process as claimed in [any one of claims 1 to 10] claim 1, wherein the mixture in step (b) is heated at a temperature in the range of from ambient temperature to boiling temperature, preferably by a reflux process.

16(Amended). A process for the preparation of a stabilized alumina catalyst support material, which process comprises the steps of:

(i) providing a precursor material comprising particles of boehmite at least partially coated with a layer comprising a compound of barium, wherein the precursor material is produced according to a process as claimed in [any one of claims 1 to 15] claim 1; and

(ii) heating the coated particles of boehmite to a temperature at which least some of the boehmite transforms to gamma-alumina.

18(Amended). A process for the preparation of an automotive catalytic converter, which process comprises the steps of:

(i) providing a precursor material comprising particles of boehmite at least partially coated with a layer comprising a compound of barium, wherein the precursor material is produced according to a process as claimed in [any one of claims 1 to 15] claim 1;

(ii) dispersing a mixture comprising said precursor material on a metal or ceramic monolithic substrate; and

(iii) heating said mixture comprising said precursor material to a temperature at which at least some of the boehmite transforms to gamma-alumina.

20(Amended). A process as claimed in claim 16 [or claim 18], wherein heating is carried out at a temperature in the range of from 500 to 1000°C.

21(Amended). A process as claimed in claim 17 [or claim 19], wherein the further heating is carried out at a temperature in the range of from 1000 to 1400°C.

22(Amended). A process for the preparation of a catalyst, such as an automotive catalytic converter, which process comprises the step of dispersing a stabilized alumina catalyst support material prepared according to the process of claim 16 [or claim 17] on a metal or ceramic monolithic substrate.

27(Amended). A stabilized alumina catalyst carrier or support material as claimed in claim 25 [or claim 26] having a specific surface area of at least 31 m²/g after exposure to a temperature of 1400°C 1 hour.

12/pstr

Stabilization of Transition Alumina

5 The present invention relates to the general field of catalysts, such as automobile exhaust catalysts, and, in particular, to a stabilized transition alumina catalyst support material.

10 Rapid developments in combustion technology coupled with the increasing demands for reduced levels of pollutants emitted by combustion processes has created a demand for new catalyst materials. In particular, new catalysts are needed for temperatures above 1000°C, not only for the removal of pollutants, for example to clean up automobile exhausts, but also
15 for catalytic combustion, for example in heaters, boilers, and gas turbines. The catalytic combustion of fuels offers the advantages of a more uniform combustion temperature and a more complete combustion of the fuel. The result is a higher thermal
20 efficiency, together with less nitrogen oxide formation through the prevention of localised high temperature areas.

25 Automobile exhaust catalysts are used to reduce automobile exhaust emissions by converting CO, unburned hydrocarbons and NO_x in the exhaust gas to an environmentally less harmful mixture comprising CO₂, H₂O and N₂. The catalysts typically comprise a metal or ceramic monolithic carrier, such as Cordierite
30 (2MgO.2Al₂O₃.5SiO₂), coated with a high surface area catalyst support material. Catalytically active noble metal particles such as, for example, Pt, Pd and/or Rh, and other additives, are dispersed on and within the support material.

- 2 -

5 The catalytically active particles need to be held in the exhaust gas stream on a support material which has a high specific surface area to maximize their exposure to the gas stream and thereby maximize the catalytic efficiency. Gamma-alumina and the closely related transition aluminas theta and delta are commonly used as support materials because of their high specific surface areas. In particular, gamma-alumina is unusual among ceramic materials in that it has a very high specific surface area, for example exceeding 100 m²/g, compared with the more typical values of from about 1 to 5 m²/g for other ceramic materials. However, gamma and the other transition aluminas are metastable phases, which are subject to thermal degradation and undergo a phase transition at temperatures exceeding about 1000°C to the thermodynamically more stable alpha-alumina (see Figures 1 and 9B) with the corollary of a significant reduction in the specific surface area; typically from about 100 m²/g to about 1 to 2 m²/g. This results in a corresponding reduction in catalytic activity. The transformation sequence for boehmite is gamma → delta → theta → alpha. Gamma, delta and theta alumina all have a cubic oxygen sub-lattice and the only difference between them is in the progressive increase in ordering of aluminium atoms on the cation deficient aluminium sub-lattice. In contrast, alpha-alumina has a hexagonal oxygen sub-lattice and its formation involves a total restructuring of the lattice.

Thermal degradation of transition alumina is a kinetic effect which accelerates rapidly with increasing temperature. Whereas at 600°C it is too slow to be significant within the lifetime of a catalytic converter, at 1100°C it may be complete

- 3 -

within 1 hour. Catalytic converters are operated at the upper limit of their stability and catalyst support materials with improved resistance to thermal degradation would allow higher operating temperatures and a longer service life.

Automotive exhaust gases from petrol engines may contain up to 3% CO and 100 ppm unburned hydrocarbons which react exothermically in the catalytic converter with NO_x to yield an environmentally less harmful mixture of gases, namely CO₂, H₂O and N₂. Peak operation temperatures of from about 700°C to about 800°C are frequently achieved. The current trend to reduce CO₂ emissions has resulted in increased use of diesel engines with lower fuel consumption and hence lower CO₂ emissions. However, diesel engines operate at higher exhaust temperatures; approximately 800°C is normal with peak temperatures of up to about 1300°C.

The conventional solution to the thermal degradation of gamma and the other transition aluminas is the use additives, such as, for example, Zr, Ba, Ce, La, Si, either singly or in combination. US-4,624,940 relates to a catalyst for the oxidation of carbon monoxide and unburned hydrocarbons in an exhaust stream. The catalyst comprises support particles of stabilized gamma-alumina having palladium dispersed thereon and can remain effective after exposure to temperatures greater than 1000°C. The gamma-alumina is stabilized by either (i) incorporation of a combination of lanthana and baria into the alumina particles, or (ii) impregnation of the alumina particles with an aqueous dispersion of a high molecular weight polyorganosiloxane followed by heating to a temperature which is effective to

- 4 -

decompose the polyorganosiloxane. The additives are introduced by soaking an alumina powder or a solid alumina precursor in a solution containing the additive followed by drying and calcining at a temperature sufficient to effect stabilization. An alternative but less widely used method of introducing additives is by a coprecipitation process as described in US-4,056,489. In this manner an intimate admixture of alumina and the stabilizing additives is formed, which is then calcined.

The present inventors have now devised a process whereby transition alumina can be further stabilized against thermal degradation to alpha-alumina. This results in a catalyst which is effective at higher operating temperatures and for longer times compared with the prior art catalysts.

Accordingly, in a first aspect the present invention provides a process for the production of a precursor material for use in the preparation of a stabilized alumina catalyst support material, which process comprises the steps of:

- (a) providing a mixture comprising an aluminium containing compound and at least one additive selected from barium, zirconium, cerium, lanthanum, calcium and silicon containing compounds; and
- (b) heating the mixture under conditions to form particles of boehmite at least partially coated with a layer comprising at least one of barium, zirconium, cerium, lanthanum, calcium and silicon and/or at least one compound containing barium,

- 5 -

zirconium, cerium, lanthanum, calcium and/or silicon.

Boehmite is a mineral form of aluminium monohydroxide (AlO.OH).

5 The additive is preferably a barium containing compound. The layer formed on the particles of boehmite preferably comprises barium carbonate.

10 An important advantage of the process according to the present invention is that the particles of boehmite can be substantially coated completely with the said layer.

15 The mixture in step (a) is preferably a substantially homogenous mixture. The substantially homogenous mixture may be formed by any process which results in an intimate admixture. For example, by forming a slurry of boehmite and/or gamma alumina with a water-soluble compound of barium, such as barium nitrate, or, alternatively, by adding a barium containing solution to a gel comprising an aluminium containing compound.

20 A particularly preferred process for forming the substantially homogenous mixture is by homogeneous precipitation, ideally homogeneous precipitation of water-soluble aluminium and barium containing compounds, for example salts thereof. Suitable aluminium compounds include aluminium nitrate, aluminium sulphate and/or aluminium chloride, preferably aluminium chloride hexahydrate ($\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$). Suitable barium compounds include barium nitrate, 25 barium sulphate and/or barium chloride, preferably barium chloride dihydrate ($\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$).

30 Precipitation from aqueous solution is widely used for producing ceramic oxide powders. 35 Precipitation occurs by adding a precipitating ligand

- 6 -

(anion) to a solution containing cations of the appropriate metal. If the precipitating ligand is added directly by simply pouring one solution into another then there is little control of the chemistry during precipitation because of the large and inhomogeneous gradients in solution concentration. Better control of chemical and morphological characteristics can be achieved if the precipitating ligands are generated "in situ", simultaneously and uniformly throughout the solution. This is known as a "homogeneous" precipitation process. In a preferred aspect of the present invention the mixture in step (a) is formed by homogeneous precipitation in which a base is generated within a metal salt solution through thermal decomposition of a suitable water-soluble reagent. Suitable water-soluble reagents include any compound that decomposes on heating in aqueous solution to release NH_3 groups, thereby forming ammonium hydroxide which increases the pH of the solution. Examples include urea and amines, such as hexamethylene tetramine. Suitable compounds of aluminium include aluminium nitrate, aluminium sulphate and/or aluminium chloride, preferably aluminium chloride hexahydrate ($\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$). Suitable compounds of barium include barium nitrate, barium sulphate and/or barium chloride, preferably barium chloride dihydrate ($\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$). A small amount of a dispersant, for example polyvinylpyrrolidone, may also be added to the metal salt solution. The dispersant improves the uniformity of the precipitation product by reducing the tendency of the precipitate particles to stick together. Additionally, diffusion of ionic species through an adsorbed dispersant layer may be the rate limiting step to further particle growth.

- 7 -

0992651-11501
5 The mixture comprising aluminium and barium
containing compounds will generally be in the form of
a gel when produced by the homogeneous precipitation
process described above, although the mixture can take
10 other forms including, for example, hydroxides,
molecular and ionic complex species or compounds. The
gel is a three-dimensional network of colloidal
particles of mainly amorphous aluminium hydroxide
containing a large amount of water (for example
15 approximately 170 g of wet gel yields approximately 7
g of dried powder) and in which the barium additive is
trapped. The consistency of the gel is generally
between that of a thick paste and a jelly, and it
easily breaks up when stirred in an excess of water
and slowly settles to the bottom again.

20 The mixture comprising aluminum and barium
containing compounds may also be formed by adding a
water-soluble barium compound, for example barium
nitrate, barium sulphate and/or barium chloride, to a
slurry of, for example, boehmite in water.
Advantageously, an organic acid, such as a carboxylic
acid is added to the mixture. A preferred carboxylic
25 acid is oxalic acid. In a particularly preferred
aspect, a solution comprising barium nitrate and
oxalic acid is added to the slurry of boehmite in
water and the mixture is then heated under conditions
to produce particles of boehmite at least partially
coated with a layer comprising barium and/or a barium
30 compound.

35 The heating in step (b) is preferably achieved by
a hydrothermal treatment, which is typically carried
out on an aqueous solution of the mixture of aluminium
and barium containing compounds in a pressure vessel,

- 8 -

such as an autoclave, at a temperature above the normal boiling point and under applied or autogenous pressure to substantially prevent boiling. The hydrothermal treatment will generally be performed at a temperature typically in the range of from 90 to 300°C, more typically from 100 to 220°C. It will be appreciated that the pressure required to substantially prevent boiling will depend on the applied temperature. Typically, the hydrothermal treatment is performed at a pressure in the range of from 1 to 150 bar, more typically from 5 to 50 bar, still more typically from 10 to 40 bar. The hydrothermal treatment is typically carried out for a time in the range of from 30 minutes to 25 hours, more typically from 1 to 10 hours.

The heating in step (b) may also be achieved by heating at a temperature in the range of from ambient temperature to boiling temperature. This may readily be achieved by a reflux operation, which may be carried out, for example, at atmospheric pressure and a temperature in the range of from 80 to 120°C, preferably approximately 100°C.

In one aspect of the present invention, the process of forming coated boehmite particles from a solution comprising aluminium and barium containing compounds can be seen to involve sequential precipitation: first, the formation of an aluminium hydroxide gel; second the precipitation of barium and/or a barium compound (probably as a carbonate) within the gel at a higher pH to give a uniform mixture; third crystallization of boehmite particles; and fourth precipitation or re-precipitation of barium and/or a barium compound (probably as a carbonate) on

- 9 -

the surface of the boehmite particles.

The product of the process according to the first aspect of the present invention may subsequently be dried and then calcined. Drying and calcination will generally be carried out in air or other oxidizing atmosphere.

In a second aspect, the present invention provides a process for the preparation of a stabilized alumina catalyst support material, which process comprises the steps of:

- (i) providing a precursor material comprising particles of boehmite at least partially coated with a layer comprising barium and/or a compound of barium; and
- (ii) heating the coated particles of boehmite to a temperature at which at least some of the boehmite transforms to gamma-alumina.

In the second aspect, heating is typically carried out at a temperature in the range of from 500 to 1000°C. The process may also include an additional step of further heating whereby at least some of the gamma-alumina transforms to theta and/or delta-alumina. The further heating is typically carried out at a temperature in the range of from 1000 to 1400°C. Heating will generally be carried out in air or other oxidizing atmosphere.

The precursor material comprising particles of boehmite may be unsupported or may be provided in the form of a washcoat on a catalyst carrier substrate,

- 10 -

such as a ceramic honeycomb.

5 The gamma-alumina particles produced according to the second aspect of the present invention may be provided in the form of a dry powder which can be redispersed in an appropriate liquid, such as water, to form a slurry or washcoat.

10 The precursor material may be produced according to a process as herein described with reference to the first aspect of the invention.

15 In a third aspect, the present invention provides a process for the preparation of a catalyst, such as an automotive catalytic converter, which process comprises the steps of dispersing a precursor mixture comprising particles of boehmite at least partially coated with a layer comprising barium and/or a compound of barium on a metal or ceramic monolithic substrate, followed by heating the precursor mixture to a temperature at which at least some of the boehmite transforms to gamma-alumina.

25 In the third aspect, heating is typically carried out at a temperature in the range of from 500 to 1000°C. The process may also include an additional step of further heating whereby at least some of the gamma-alumina transforms to delta and/or theta-alumina. The further heating is typically carried out at a temperature in the range of from 1000 to 1400°C. Again, heating will generally be carried out in air or other oxidizing atmosphere.

35 The mixture can be dispersed on to metal or ceramic monolithic substrate in a conventional manner,

- 11 -

for example by dipping the substrate one or more times into an aqueous slurry (washcoat) of the coated boehmite powder, followed by drying.

5 In the second and third aspects of the present invention, the particles of gamma-alumina, theta, delta-alumina and/or boehmite will typically be at least partially coated with a layer comprising barium and/or a compound of barium. The layer will typically
10 comprise one or more of barium carbonate, barium oxide and/or barium aluminate depending on the heat-treatment. For gamma-alumina particles, the layer will generally comprise barium oxide as the major constituent, whilst for theta and delta-alumina
15 particles, the layer will generally comprise one or more barium aluminates as the major constituent(s).

In a fourth aspect, the present invention provides a process for the preparation of a catalyst, such as an automotive catalytic converter, which
20 process comprises the step of dispersing a stabilized alumina catalyst support material as herein described on a metal or ceramic monolithic substrate.

25 Again, the stabilized alumina catalyst support material may be dispersed on to the metal or ceramic monolithic substrate in a conventional manner, for example by dipping the substrate one or more times into an aqueous slurry (washcoat) of stabilized
30 alumina powder, followed by drying.

In a fifth aspect, the present invention provides a precursor material for use in the preparation of catalyst support material, which precursor material
35 comprises particles of boehmite and/or transition

- 12 -

alumina at least partially coated with a layer comprising barium and/or a compound of barium.

Each particle of boehmite and/or transition alumina preferably has a substantially uniform surface layer comprising barium and/or a compound of barium. The layer will typically comprise one or more of barium carbonate, barium oxide and/or barium aluminate depending on the heat-treatment used to process the particles. For boehmite, the layer will generally comprise barium carbonate as the major constituent; for gamma-alumina particles, the layer will generally comprise barium oxide as the major constituent; and for theta and delta-alumina particles, the layer will generally comprise one or more barium aluminates as the major constituent(s).

The size of the particles may range from 10 nm to 5 μm , but the particles typically have an average particle size in the range of from 100 to 500 nm. In general the particles take the form of thin prismatic crystals with an average length to width ratio typically of about 3:1. The average length to thickness ratio is typically in the range of from 10:1 to 20:1. The crystal sizes are for the longest dimension. The surface layer comprising barium and/or a compound of barium will typically be up to 5 nm in thickness, more typically up to 3 nm in thickness.

The precursor material may provided in the form of an aqueous slurry, a non-aqueous slurry or a powder.

In a sixth aspect, the present invention provides a stabilized alumina catalyst support material which

- 13 -

comprises gamma, theta and/or delta-alumina, and which withstands substantial degradation to alpha-alumina after exposure to a temperature of approximately 1400°C for 1 hour when heated in, for example, air or other oxidizing atmosphere.

In this aspect, the stabilized alumina catalyst support material preferably has a specific surface area of at least 20 m²/g, more preferably at least 25 m²/g, still more preferably at least 31 m²/g, after exposure to a temperature of 1400°C for 1 hour. The specific surface area is measured using the standard BET method.

In a seventh aspect, the present invention provides a catalyst, such as an automotive catalytic converter, which comprises a metal or ceramic monolithic substrate having dispersed thereon a stabilized alumina catalyst support material as herein described.

It will be appreciated that one or more catalytically active noble metal particles, for example Pt, Pd and/or Rh, optionally together with other additives, will typically be dispersed on the support material.

In the third, fourth and seventh aspects the catalyst may be a 3-way automotive catalytic converter.

The monolithic substrate can be formed from any of the conventional ceramic materials in the art including, for example, one or more of cordierite, cordierite-alpha alumina, silicon nitride, silicon

- 14 -

carbide, mullite, zirconia-mullite, spodumene,
alumina-silica-magnesia and/or zirconium silicate.
Examples of suitable metal substrates include
austenitic stainless steel and high aluminium ferritic
stainless steel.

It will be appreciated that the stabilized
alumina according to the present invention may be used
in conjunction with other components conventionally
used in the field of catalysis, such as, for example:
catalytic promoters, including Ni, Co, Mn, Cr_2O_3 , TiO_2 ,
 ZrO_2 and/or CeO_2 ; washcoat stabilization oxides,
including cordierite, mullite and/or magnesium
aluminium titanate; additional stabilization
additives, including Zr, Ce, La, Ca, Cu, Th and/or Si.

The present invention will now be described
further, by way of example, with reference to the
following drawings, in which:

Figure 1 is a TEM micrograph of alumina without
additive showing the transformation of
transition-alumina to alpha-alumina with a
reduction in specific surface area;

Figure 2 is a TEM micrograph of a hydrothermally
treated material showing large alpha-alumina
crystals and nanoparticles of barium aluminate.
The material was formed by heating a
hydrothermally treated alumina-baria precursor to
1400°C for 1 hour in argon;

Figure 3 is a TEM micrograph showing a porous
boehmite crystal pseudomorph consisting of gamma-
alumina coated with a barium oxide layer. The

- 15 -

material was formed by heating a hydrothermally-treated precipitate of an alumina-baria precursor to 1000°C for 1 hour;

5 Figures 4 shows TEM micrographs of boehmite pseudomorphs consisting of theta-alumina particles exhibiting a twinned structure and some residual porosity. The material was formed by heating a hydrothermally-treated precipitate of alumina-baria precursor to 1400°C for 1 hour;

10 Figure 5 A and B are TEM micrographs showing a comparison between boehmite produced by a hydrothermal treatment of a precipitate of an alumina-baria precursor and theta-alumina formed by heating the same boehmite: (A) boehmite produced by a hydrothermal treatment of a precipitate of an alumina-baria precursor, and (B) theta-alumina formed by heating the boehmite to 1400°C for 1 hour;

15 Figure 6 is a flow chart showing a synthesis procedure for making alumina-baria precursor gels;

20 Figure 7 is a DTA analysis of an autoclave treated alumina-baria precursor gel for various heating rates.

25 Figures 8 A-D are TEM micrographs of an oven-dried alumina-baria precursor gel: (A) and (B) microcrystalline bayerite before calcination, bright and dark filed respectively, (C) and (D) gamma-alumina with crystals of β_{11} after 1 hour at 1000°C, bright and dark field respectively;

- 16 -

Figures 9 A-F are TEM micrographs of a microwave treated alumina-baria precursor gel: (A) morphology after microwave drying, (B) large alpha-alumina crystal growing in matrix of theta-alumina after 1 hour at 1200°C in air, (C) and (D) nano-particles of β_1 inside alpha-alumina crystals after 1 hour at 1400°C in air, (E) and (F) nano-particles of β_1 inside and on the surface of alpha-alumina crystals after 1 hour at 1400°C in argon;

Figure 10 A-C are TEM micrographs of an autoclave treated alumina-baria precursor gel: (A) well-developed boehmite crystals after hydrothermal treatment of 4 hours at 210°C, (B) porous boehmite pseudomorphs after 1 hour at 1000°C, and (C) detail showing size, shape and orientation of pores;

Figures 11 A-D are TEM micrographs of an autoclave treated alumina-baria precursor gel after calcination for 1 hour at 1400°C in air: (A) boehmite pseudomorphs comprising theta-alumina, (B) finely twinned theta-alumina with equiaxed pores, (C) and (D) details of twinned theta-alumina structure showing twin lamella as narrow as 5 nm; and

Figures 12 A and B are TEM micrographs of an autoclave treated alumina-baria precursor gel after calcination for 1 hour at 1400°C in argon: (A) nano-particles of $\text{BaAl}_{12}\text{O}_{19}$ in a matrix of theta-alumina transformed extensively to alpha-alumina, (B) detail showing $\text{BaAl}_{12}\text{O}_{19}$ particles predominantly on the surface of the alpha-

- 17 -

alumina.

The present inventors have investigated a number of ways of introducing barium additives to gamma-
alumina and have found that they all yield similar
results to those already known except for two methods
which yielded a product with a markedly superior
resistance to thermal degradation. The process
according to the present invention relies on a
hydrothermal or refluxing treatment of a substantially
homogeneous mixture of aluminium and barium containing
compounds.

The hydrothermal or refluxing treatment results
in crystals of boehmite which appear to be coated with
a layer of barium or a barium compound some nanometres
thick. The layer is thought to be barium carbonate.
Boehmite (AlO.OH) is an aluminium monohydroxide which
is a gamma-alumina precursor. On heating to about
500°C, boehmite transforms to gamma-alumina without a
substantial change in external crystal shape, but with
the development of internal porosity, which is thought
to be responsible for the unusually high specific
surface area of gamma-alumina. It seems that by
depositing the barium as a substantially uniform
surface coating, the effectiveness of the barium as a
stabilizer for gamma-alumina is maximized and that the
hydrothermal or refluxing method is therefore markedly
superior to other methods such as calcining that yield
a more or less homogeneous mixture of phases.

On heating, the hydrothermally treated material
transforms to gamma-alumina at temperatures up to
about 1000°C (see Figure 3) and transition alumina
(theta and delta phases) at temperatures up to about

- 18 -

1400°C for about 1 hour (see Figure 4). The layer comprising the barium-containing compound, possibly barium carbonate, transforms first to barium oxide on heating to about 830°C in the pure state and then subsequently to barium aluminate.

After heating to a higher temperature, X-ray diffraction showed that the barium formed barium oxide at temperatures up to about 1000°C and various barium aluminates (for example BaAl_2O_4 , $\text{BaAl}_{13.2}\text{O}_{20.8}$ and $\text{BaAl}_{9.2}\text{O}_{14.8}$) at higher temperatures. These phases were present in all samples, not just in the case of the hydrothermally treated samples.

Barium aluminates are believed to be effective in preventing the nucleation of alpha-alumina if a sufficient quantity is present on the surface of gamma-alumina particles to fill surface defects which otherwise act as nucleation sites for the phase transformation to alpha-alumina.

In microwave-treated materials after about 1 hour at approximately 1200°C, barium aluminate ($\text{BaAl}_{13.2}\text{O}_{20.8}$) was visible as nano-sized particles predominantly inside much larger alpha-alumina crystals in transmission electron micrographs and was particularly well-developed after about 1 hour at approximately 1400°C (see Figures 9C and 9D), whereas in hydrothermally-treated alumina such particles were absent after treatment at 1400°C, except when heated in argon instead of air, in which case nano-sized particles of barium aluminate ($\text{BaAl}_{12}\text{O}_{19}$) were visible, predominantly on the surface of the alpha-alumina crystals (see Figure 2). The conclusion is that barium aluminate was present as a thin surface layer

- 19 -

in the hydrothermally treated material even after 1 hour at 1400°C in air, but not in a similar starting material treated by oven-drying or microwave-drying because those materials transformed or partly transformed to alpha-alumina after 1 hour at 1200°C, a temperature similar to that reported by other authors. An important step therefore seems to be the formation of a generally uniform surface layer of barium and/or a barium compound on the boehmite crystals during the hydrothermal or refluxing treatment.

The length of hydrothermal treatment is important in determining the specific surface area of the product after the subsequent high temperature treatment. After heating for 1 hour at 1400°C, a material that was previously given a 4 hour hydrothermal treatment at approximately 210°C resulted in theta-alumina with a specific surface area of 6 m²/g. In contrast, after heating for 1 hour at 1400°C, a material that was previously given a 22 hour hydrothermal treatment at 210°C resulted in theta-alumina with a more useful specific surface area of 31 m²/g. Such a material would retain sufficient surface area to remain effective as a catalyst support material. Figures 5A and 5B shows that there is remarkably little change in general appearance of the hydrothermally treated material before (5A) and after (5B) heating for 1 hour at 1400°C which demonstrates the effectiveness of this material in resisting thermal degradation.

The present invention will now be described further with reference to the following Examples and Comparative Examples.

Examples

Aluminium chloride hexahydrate ($\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$) (Merck, crystalline, purified), barium chloride dihydrate ($\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$) (Baker, analysed, reagent grade), urea (CON_2H_4) (Acros, p.a.) and polyvinylpyrrolidone (PVP, Aldrich, K-30, special grade) were used as received. Solutions of the above reagents were filtered before use.

Alumina and alumina doped with approximately 10 mol% baria were synthesised by homogeneous precipitation from a metal salt solution induced by the thermal decomposition of urea as shown in the flow chart in Figure 6. On heating the solution to approximately 86°C , the urea decomposed to produce a controlled and uniform increase in pH until precipitation occurred. Very fine sol particles sedimented from the solution after standing for several hours at room temperature. The reaction product was centrifuged to separate the characteristically voluminous and transparent gel. The gel was washed twice with double its own volume of distilled water to remove most of the anion and organic impurities. About 1200 ml of the transparent gel was produced from 5 l of solution in a typical experiment.

Some Ba^{2+} was detected qualitatively in the washing water by precipitation with SO_4^{2-} . The loss of barium could be avoided by omitting the washing step, in which case the chloride anion was removed by sublimation of NH_4Cl , either during microwave drying, or during calcination in the case of the oven-dried specimens. For autoclave treated gels, washing

- 21 -

carried out after crystallisation of boehmite and barium carbonate removed chloride ions without a loss of barium.

5 The alumina and alumina-baria gels were treated in a variety of different ways: air-dried at room temperature; oven-dried at a temperature in the range of from 90 to 100°C; dried in a microwave oven; and hydrothermally treated in an autoclave. The air-dried
10 gel was prepared by leaving the gel in a glass beaker at room temperature until a solid formed which was then crushed in an agate mortar to produce a fine powder. The oven-dried gel was prepared by heating the gel in air in an oven at about 90 to 100°C for
15 approximately 24 hours; the solid was then crushed as above to produce a fine powder. The microwave-dried samples of gel were prepared in a standard domestic microwave oven at a power setting of 500 W using cycles of 5 min power on and from 2 to 3 min power
20 off, because of extensive foaming of the gel due to the release of gaseous products. The dried solid was crushed as above and later calcined at different temperatures. The hydrothermal treatment was conducted on 150 ml samples of centrifuged gel in 230
25 ml Teflon-lined pressure-tight steel containers. A heating rate of about 0.5°C/min to a temperature of $210 \pm 5^\circ\text{C}$ was used. The hold time at the temperature was about 4 hours. The resulting brownish liquor with a pH in the range of from 8 to 8.5 was decanted and
30 the crystallised product was dried at a temperature from 90 to 100°C. The crystalline product was calcined in either air or argon.

35 The size of the colloidal particles in the alumina and alumina-baria gels was measured by Small

- 22 -

Angle X-ray Scattering (SAXS); evaluation of the results of this technique was based on the assumption that the gel was a fractal system (see J. Teixeira, *J. Appl. Cryst.* 21 (1988) 781). The phase composition of the samples was determined by XRD, either in-situ during calcination (Siemens D5000), or conventionally before and after calcination at various temperatures (Philips PW173). Thermal decomposition of the gel to the calcined products on heating to 1400°C was observed by differential thermal analysis and thermogravimetry (DTA/TG, Netsch Simultaneous Analyser STA 409). Transmission electron microscopy (TEM, Philips EM 400) and High Resolution Transmission Electron Microscopy (HRTEM, JEOL 4000 EX/II with a point resolution of 0.165 nm at 400 kV) were used for analysis of the fine structure of uncalcined and calcined samples. Electron microscopy samples were prepared by dispersing the powder in ethyl alcohol, placing a drop of this suspension on a carbon-coated copper grid, and allowing the alcohol to evaporate. The specific surface area was measured by BET nitrogen adsorption (Coulter-Omnisorp 360).

The important characteristic of the homogeneous precipitation process was that it yielded a product substantially uniform both in chemical composition and particle size. The decomposition of urea on heating the solution released ammonium ions into solution in a very uniform manner resulting in a gradual and uniform increase in pH. These uniform conditions ensured that the alumina-baria composite gel which precipitated from the dissolved chloride salts was substantially chemically homogeneous on a colloidal scale. The precipitation kinetics were easily controlled through the heating rate and the ageing temperature. The flow

- 23 -

chart in Figure 6 summarises the steps and conditions of the synthesis.

The size of the sol particles in the gel samples was calculated from SAXS measurements as shown in Table 1. The alumina and alumina-baria precursor gels both consisted of particles of about 4 nm diameter and the fractal dimension indicates that both gels were in a similar state of aggregation.

	Alumina	Alumina-baria
Fractal dimension	2.2	2.4
Particle diameter (nm)	3.7	4.1

Table 1: Properties of alumina and alumina-baria precursor gel measured by SAXS

The phase evolution during calcination was studied by X-ray diffraction "in-situ" and conventionally, i.e. on specimens measured at room temperature after calcination. The main results for the conventional measurements are summarised in Table 2 with phases listed in order of abundance. The phases in alumina-baria samples were compared with those in pure alumina samples to obtain information on the effects of each of the gel treatment techniques.

Sample	Phases determined by XRD				
	air				argon
	uncalcined	1000°C 1h	1200°C 1h	1400°C 1h	1400°C 1h
oven dried 100°C 24h	amorphous bayerite urea (tr)	γ -Al ₂ O ₃ BaO	α -Al ₂ O ₃ BaAl ₂ O ₄ β_I	α -Al ₂ O ₃ β_I	nd
microwave dried	amorphous boehmite BaAl ₁₂ O ₁₉ (i)	γ -Al ₂ O ₃ β_{II} (tr)	α -Al ₂ O ₃ θ -Al ₂ O ₃ β_I	α -Al ₂ O ₃ β_I	α -Al ₂ O ₃ β_I
autoclaved 4h	boehmite BaCO ₃ (i)	γ -Al ₂ O ₃ δ -Al ₂ O ₃ BaO(i)	θ -Al ₂ O ₃ δ -Al ₂ O ₃ BaAl ₂ O ₄	θ -Al ₂ O ₃ δ -Al ₂ O ₃ BaAl ₂ O ₄ β_{II}	α -Al ₂ O ₃ θ -Al ₂ O ₃ BaAl ₁₂ O ₁₉
autoclaved 22h	nd	nd	nd	θ -Al ₂ O ₃ BaAl ₂ O ₄ β_{II}	nd

$\beta_I = \text{BaAl}_{13.2}\text{O}_{20.8}$, ($\text{BaO} \cdot 6.6\text{Al}_2\text{O}_3$), $\beta_{II} = \text{BaAl}_{9.2}\text{O}_{14.8}$,
 ($\text{BaO} \cdot 4.6\text{Al}_2\text{O}_3$), tr = traces, (i) only most intense
 peaks visible, nd = not determined

Table 2: Phase composition of alumina-baria samples
 before and after calcination of dehydrated gel.

The alumina-baria gel dried at room temperature appeared to be amorphous from the XRD spectra. The only indication of the presence of a crystalline phase was a small amount of retained urea. The XRD spectrum for the oven-dried sample (100°C) also indicated predominantly amorphous material and the presence of very poorly crystallized bayerite Al(OH)₃, besides traces of crystalline urea. There was no indication of the presence of any crystalline barium containing

phase.

In the microwave-treated alumina sample, "in situ" XRD indicated the presence of microcrystalline boehmite to approximately 400°C at which temperature the sample became amorphous and remained amorphous to about 800-850°C when transition alumina (γ - Al_2O_3) appeared. At about 950-1000°C transformation of γ - Al_2O_3 to α - Al_2O_3 began and was almost complete at 1200°C. In the microwave-treated alumina-baria sample, the phase initially present was also microcrystalline boehmite (or pseudo boehmite) which became amorphous at about 400°C. The transition aluminas that formed were γ - Al_2O_3 and θ - Al_2O_3 instead of just γ - Al_2O_3 as in the pure alumina sample. The transformation from θ - Al_2O_3 to α - Al_2O_3 began at about 1050-1100°C. After calcination at 1000°C for 1 hour traces of the $\text{BaAl}_{9.2}\text{O}_{14.8}$ (β_{II}) phase were present. After calcination for 1 hour at 1200°C or 1400°C, $\text{BaAl}_{13.2}\text{O}_{20.8}$ (β_I) was present instead of β_{II} . Calcination in argon or argon and air for 1 hour at 1400°C gave the same result.

In the autoclave treated alumina and alumina - baria precursor gels XRD showed that the samples initially consisted of well-crystallised boehmite which decomposed at 450-500°C to an amorphous phase. Two rather weak peaks corresponding to the main peaks in the spectrum for γ -barium carbonate were also detected. In-situ XRD showed that the two BaCO_3 peaks persisted to 700°C at which temperature a peak corresponding to the main BaO peak appeared and transition alumina formed. After 1 hour at 1000°C, γ - Al_2O_3 , δ - Al_2O_3 and BaO were the phases present. After 1 hour at 1200°C γ - Al_2O_3 , δ - Al_2O_3 and BaAl_2O_4 were

present. After 1 hour at 1400°C, θ -Al₂O₃ and δ -Al₂O₃ were present with BaAl₂O₄ and BaAl_{9.2}O_{14.8} (β_{II}). A trace amount of α -Al₂O₃ was found in this sample calcined in air for 1 hour and, after 10 hours at 1400°C, θ -Al₂O₃ was still the predominant phase. However, in samples calcined in argon for 1 hour at 1400°C the principle phase was α -Al₂O₃.

The DTA/TG curves in Figure 7 for the hydrothermally treated baria-doped samples show well-defined exothermic peaks corresponding to the decomposition of boehmite by dehydroxylation. A displacement of the dehydroxylation peaks to higher temperatures with increased heating rate indicates the slow kinetics of dehydration. The theoretical weight loss for the dehydroxylation of boehmite is 15%. In Ba doped boehmite a weight loss of less than 15% was expected whatever barium compound was present, for example carbonate or hydroxide, due to the high atomic weight of barium compared with aluminium. There was no indication of Ba phases such as BaCO₃ or Ba(OH)₂·xH₂O which have a number of well-defined peaks by which they could easily be identified if present.

BET surface area measurements were made on samples calcined for 1 hour at either 1200°C or 1400°C. For calcination at 1200°C, a value of 45.2 m²/g was measured for specimens autoclaved for 4 hours. For calcination at 1400°C, values of 5.9 m²/g and 31.2 m²/g were obtained for specimens autoclaved for 4 hours and 22 hours respectively. The longer autoclave treatment resulted in larger boehmite crystal size in the sample before calcination. After calcination at 1400°C, the phase composition of the specimens was also different. From XRD analysis, the

- 27 -

phases present in order of abundance in the specimen autoclaved for 4h were: θ - Al_2O_3 , δ - Al_2O_3 , BaAl_2O_4 and β_{II} . The specimen autoclaved for 22 hours differed in that δ - Al_2O_3 was absent.

5

TEM examination showed that considerable morphological differences existed between oven-dried, microwave-dried and hydrothermally treated samples, both before and after calcination.

10

Before calcination, the oven-dried material consisted of densely agglomerated fine crystals, Figures 8A and 8B, which were identified by electron diffraction as bayerite. The microwaved material was much less tightly agglomerated as shown in Figure 9A and consisted of boehmite crystallites of about 5 nm as seen by HRTEM (not shown here) which approximately correspond to the size of the colloidal particles in the gel. The hydrothermally treated sample consisted of plate-like boehmite crystals about 250 x 100 x 10 nm as shown in Figure 10A. There was no evidence of a separate Ba containing phase in any of the uncalcined samples.

15

20

25

TEM micrographs of oven-dried material calcined for 1 hour at 1000°C indicated the presence of a γ - Al_2O_3 matrix and finely distributed BaO (small dark particles) (see Figures 8C and 8D). The crystallite size of the γ - Al_2O_3 and barium containing phase was about the same after calcination as that of bayerite before calcination (see Figures 8B and 8D). After calcination at 1000°C the autoclaved material consisted of γ - Al_2O_3 which retained the shape of the boehmite crystals (pseudomorphs) as shown in Figure 10B. The pseudomorphs were porous and the pore facets

30

35

00926051-11501

- 28 -

can be seen in Figure 10C.

5 The TEM images of microwaved material calcined for 1 hour at 1200°C, Figures 1 and 9B, show the start of the transition from θ -Al₂O₃ to α -Al₂O₃ in pure alumina and in alumina-baria samples, respectively. After calcination at 1400°C for 1 hour in air or argon, the microwaved material consisted of BaAl_{13.2}O_{20.8} (β_1) particles of from 10 to 15 nm diameter and much larger α -Al₂O₃ crystals as shown in Figures 9C to 9F. The β_1 particles were predominantly within the α -Al₂O₃ crystals after calcination in air, but were distributed both within and on the surface of the α -Al₂O₃ crystals after calcination in argon.

15 The autoclave treated material, calcined at 1400°C in air, consisted mainly of finely twinned θ -Al₂O₃ with twin planes spaced about 5 nm apart as shown in Figures 11 A-D. The boehmite pseudomorph crystal shape was still retained although the internal porosity was much less than at 1000°C, and was more spherical in shape. The θ -Al₂O₃ twin planes were parallel to the long axis of the pseudomorphs, i.e. the boehmite $\langle 001 \rangle$ direction. There was no evidence of sintering between the crystals and their size was similar to that at 1200°C. Small amounts α -Al₂O₃ were observed. Detailed HRTEM examination of transition alumina structures in the autoclave treated samples indicated that the fine twinning in θ -Al₂O₃ seen in the material calcined at 1400°C probably resulted from the ordering of cation vacancies in δ -Al₂O₃ at lower temperatures. In contrast, autoclave treated material, calcined at 1400°C in argon, consisted of large crystals of α -Al₂O₃ decorated on the surface with 10 nm particles of BaAl₁₂O₁₉ which has a

- 29 -

magnetoplumbite structure rather than a β type structure (see Figures 12A and 12B).

5 The present invention provides a process whereby
transition alumina can be further stabilized against
thermal degradation to alpha-alumina. This results in
a catalyst which is effective at higher operating
temperatures and for longer times compared with the
prior art catalysts. An additional advantage of the
10 present invention is that it is very effective at
forming thin (approximately 1 to 5 μm) washcoats of
gamma-alumina on cordierite or other ceramic
substrates. This is important because the use in the
future of ceramic honeycomb substrates with a very
15 high channel density will require the use of thinner
washcoats. The process according to the present
invention can produce thinner films where (coated)
boehmite is used in the slurry, rather than gamma-
alumina.
20

25

30

35

- 30 -

Claims:

1. A process for the production of a precursor material for use in the preparation of a stabilized alumina catalyst support material, which process comprises the steps of:

(a) providing a mixture comprising aluminium and barium containing compounds; and

(b) heating the mixture under conditions to form particles of boehmite at least partially coated with a layer comprising barium and/or a compound of barium.

2. A process as claimed in claim 1, wherein the mixture comprising aluminium and barium containing compounds is a substantially homogenous mixture.

3. A process as claimed in claim 1 or claim 2, wherein the layer formed on the particles of boehmite comprises barium carbonate.

4. A process as claimed in any one of the preceding claims, wherein the aluminium containing compound comprises aluminium chloride, aluminium nitrate, aluminium sulphate, alumina and/or boehmite.

5. A process as claimed in any one of the preceding claims, wherein the barium containing compound comprises barium chloride, barium sulphate and/or barium nitrate.

6. A process as claimed in any one of the preceding claims, wherein the mixture comprising aluminium and

- 31 -

barium containing compounds is in the form of a substantially homogeneous precipitate.

5 7. A process as claimed in claim 6, wherein the substantially homogeneous precipitate is formed by a process involving homogeneous precipitation of aluminium and barium containing compounds from salt solution.

10 8. A process as claimed in claim 7, wherein the homogeneous precipitation involves generating a base within the salt solution through thermal decomposition of a water-soluble reagent.

15 9. A process as claimed in claim 8, wherein the water-soluble reagent comprises urea and/or hexamethylene tetramine.

20 10. A process as claimed in any one of claims 7 to 9, wherein the salt solution further includes a dispersant, such as polyvinylpyrrolidone.

25 11. A process as claimed in any one of claims 1 to 5, wherein the mixture in step (a) is formed by adding a solution comprising oxalic acid and a water soluble barium compound, preferably barium nitrate, to an aqueous slurry comprising boehmite.

30 12. A process as claimed in any one of the preceding claims, wherein heating of the mixture in step (b) is achieved by a hydrothermal process.

35 13. A process as claimed in claim 12, wherein the hydrothermal treatment is carried out at a temperature in the range of from 90 to 300°C, more preferably from

- 32 -

100 to 220°C.

14. A process as claimed in claim 12 or claim 13,
wherein the hydrothermal treatment is carried out at a
pressure in the range of from 1 to 150 bar, preferably
from 5 to 50 bar.

15. A process as claimed in any one of claim 12 to
14, wherein the hydrothermal treatment is carried out
for a time in the range of from 30 minutes to 25
hours, preferably from 1 to 10 hours.

16. A process as claimed in any one of claims 1 to
11, wherein the mixture in step (b) is heated at a
temperature in the range of from ambient temperature
to boiling temperature, preferably by a reflux
process.

17. A process for the preparation of a stabilized
alumina catalyst support material, which process
comprises the steps of:

(i) providing a precursor material comprising
particles of boehmite at least partially coated
with a layer comprising barium and/or a compound
of barium; and

(ii) heating the coated particles of boehmite to a
temperature at which at least some of the
boehmite transforms to gamma-alumina.

18. A process as claimed in claim 17, comprising the
step of further heating whereby at least some of the
gamma-alumina transforms to theta and/or delta-
alumina.

- 33 -

19. A process as claimed in claim 17 or claim 18, wherein said precursor material is produced according to a process as claimed in any one of claim 1 to 16.

5 20. A process for the preparation of an automotive catalytic converter, which process comprises the steps of dispersing a mixture comprising particles of boehmite at least partially coated with a layer comprising barium and/or a compound of barium on a
10 metal or ceramic monolithic substrate, followed by heating the precursor mixture to a temperature at which at least some of the boehmite transforms to gamma-alumina.

15 21. A process as claimed in claim 20, comprising the step of further heating whereby at least some of the gamma-alumina transforms to theta and/or delta-alumina.

20 22. A process as claimed in claim 17 or claim 20, wherein heating is carried out at a temperature in the range of from 500 to 1000°C.

25 23. A process as claimed in claim 18 or claim 21, wherein the further heating is carried out at a temperature in the range of from 1000 to 1400°C.

30 24. A process for the preparation of a catalyst, such as an automotive catalytic converter, which process comprises the step of dispersing a stabilized alumina catalyst support material prepared according to the process of any one of claims 17 to 19 on a metal or ceramic monolithic substrate.

35 25. A precursor material for use in the preparation

- 34 -

of catalyst support material, which precursor material comprises particles of boehmite and/or transition alumina at least partially coated with a layer comprising barium and/or a compound of barium.

26. A precursor material as claimed in claim 25, wherein the layer formed on the particles of boehmite and/or transition alumina comprises barium carbonate, barium oxide and/or barium aluminate.

27. A precursor material as claimed in claim 25 or claim 26 provided in the form of a slurry or washcoat.

28. A stabilized alumina catalyst support material which comprises gamma, theta and/or delta-alumina, and which withstands substantial degradation to alpha-alumina after exposure to a temperature of approximately 1400°C for 1 hour.

29. A stabilized alumina catalyst support material as claimed in claim 28, wherein particles of gamma, theta and/or delta-alumina are at least partially coated with a layer comprising barium carbonate, barium oxide and/or barium aluminate.

30. A stabilized alumina catalyst carrier or support material as claimed in claim 28 or claim 29 having a specific surface area of at least 20 m²/g after exposure to a temperature of 1400°C for 1 hour.

31. A stabilized alumina catalyst carrier or support material as claimed in any one of claims 28 to 30 having a specific surface area of at least 31 m²/g after exposure to a temperature of 1400°C for 1 hour.

- 30 -

CLAIMS:

1. A process for the production of a precursor material for use in the preparation of stabilized alumina catalyst support material, which process comprises the steps of:
- (a) providing a substantially homogeneous mixture comprising aluminium and barium containing compounds; and
- (b) heating the mixture under conditions to form particles of boehmite at least partially coated with a layer comprising a compound of barium.
2. A process as claimed in claim 1, wherein the layer formed on the particles of boehmite comprises barium carbonate.
3. A process as claimed in claim 1 or claim 2, wherein the aluminium containing compound comprises aluminium chloride, aluminium nitrate, aluminium sulphate, alumina and/or boehmite.
4. A process as claimed in any one of the preceding claims, wherein the barium containing compound comprises barium chloride, barium sulphate and/or barium nitrate.
5. A process as claimed in any one of the preceding claims, wherein the mixture comprising aluminium and barium containing compounds is in the form of a substantially homogeneous precipitate.
6. A process as claimed in claim 5, wherein the substantially homogeneous precipitate is formed by a process involving homogeneous precipitation of

- 31 -

aluminium and barium containing compounds from salt solution.

5 7. A process as claimed in claim 6, wherein the homogeneous precipitation involves generating a base within the salt solution through thermal decomposition of a water-soluble reagent.

10 8. A process as claimed in claim 7, wherein the water-soluble reagent comprises urea and/or hexamethylene tetramine.

15 9. A process as claimed in any one of claims 6 to 8, wherein the salt solution further includes a dispersant, such as polyvinylpyrrolidone.

20 10. A process as claimed in any one of claims 1 to 4, wherein the mixture in step (a) is formed by adding a solution comprising oxalic acid and a water soluble barium compound, preferably barium nitrate, to an aqueous slurry comprising boehmite.

25 11. A process as claimed in any one of the preceding claims, wherein heating of the mixture in step (b) is achieved by a hydrothermal process.

30 12. A process as claimed in claim 11, wherein the hydrothermal treatment is carried out at a temperature in the range of from 90 to 300°C, more preferably from 100 to 220°C.

35 13. A process as claimed in claim 11 or claim 12, wherein the hydrothermal treatment is carried out at a pressure in the range of from 1 to 150 bar, preferably from 5 to 50 bar.

14. A process as claimed in any one of claims 11 to

- 32 -

13, wherein the hydrothermal treatment is carried out for a time in the range of from 30 minutes to 25 hours, preferably from 1 to 10 hours.

5 15. A process as claimed in any one of claims 1 to 10, wherein the mixture in step (b) is heated at a temperature in the range of from ambient temperature to boiling temperature, preferably by a reflux process.

10

16. A process for the preparation of a stabilized alumina catalyst support material, which process comprises the steps of:

15 (i) providing a precursor material comprising particles of boehmite at least partially coated with a layer comprising a compound of barium, wherein the precursor material is produced according to a process as claimed
20 in any one of claims 1 to 15; and

(ii) heating the coated particles of boehmite to a temperature at which least some of the boehmite transforms to gamma-alumina.

25

17. A process as claimed in claim 16, comprising the step of further heating whereby at least some of the gamma-alumina transforms to theta and/or delta-alumina.

30

18. A process for the preparation of an automotive catalytic converter, which process comprises the steps of:

35

(i) providing a precursor material comprising particles of boehmite at least partially coated with a layer comprising a compound of barium, wherein the precursor material is

- 33 -

produced according to a process as claimed
in any one of claims 1 to 15;

(ii) dispersing a mixture comprising said
precursor material on a metal or ceramic
monolithic substrate; and

(iii) heating said mixture comprising said
precursor material to a temperature at which
at least some of the boehmite transforms to
gamma-alumina.

19. A process as claimed in claim 18, comprising the
step of further heating whereby at least some of the
gamma-alumina transforms to theta and/or delta-
alumina.

20. A process as claimed in claim 16 or claim 18,
wherein heating is carried out at a temperature in the
range of from 500 to 1000°C.

21. A process as claimed in claim 17 or claim 19,
wherein the further heating is carried out at a
temperature in the range of from 1000 to 1400°C.

22. A process for the preparation of a catalyst, such
as an automotive catalytic converter, which process
comprises the step of dispersing a stabilized alumina
catalyst support material prepared according to the
process of claim 16 or claim 17 on a metal or ceramic
monolithic substrate.

23. A precursor material for the use in the
preparation of catalyst support material, which
precursor material comprises particles of boehmite
and/or transition alumina substantially completely
coated with a layer comprising barium carbonate,

- 34 -

barium oxide and/or barium aluminate.

24. A precursor material as claimed in claim 23 provided in the form of a slurry or washcoat.

5

25. A stabilized alumina catalyst support material which withstands substantial degradation to alpha-alumina after exposure to a temperature of approximately 1400°C for 1 hour and which comprises particles of gamma, theta and/or delta-alumina substantially completely coated with a layer comprising barium carbonate, barium oxide and/or barium aluminate.

10

15

26. A stabilized alumina catalyst carrier or support material as claimed in claim 25 having a specific area of at least 20 m²/g after exposure to a temperature of 1400°C for 1 hour.

20

27. A stabilized alumina catalyst carrier or support material as claimed in claim 25 or claim 26 having a specific surface area of at least 31 m²/g after exposure to a temperature of 1400°C 1 hour.

FOSTT 15092660

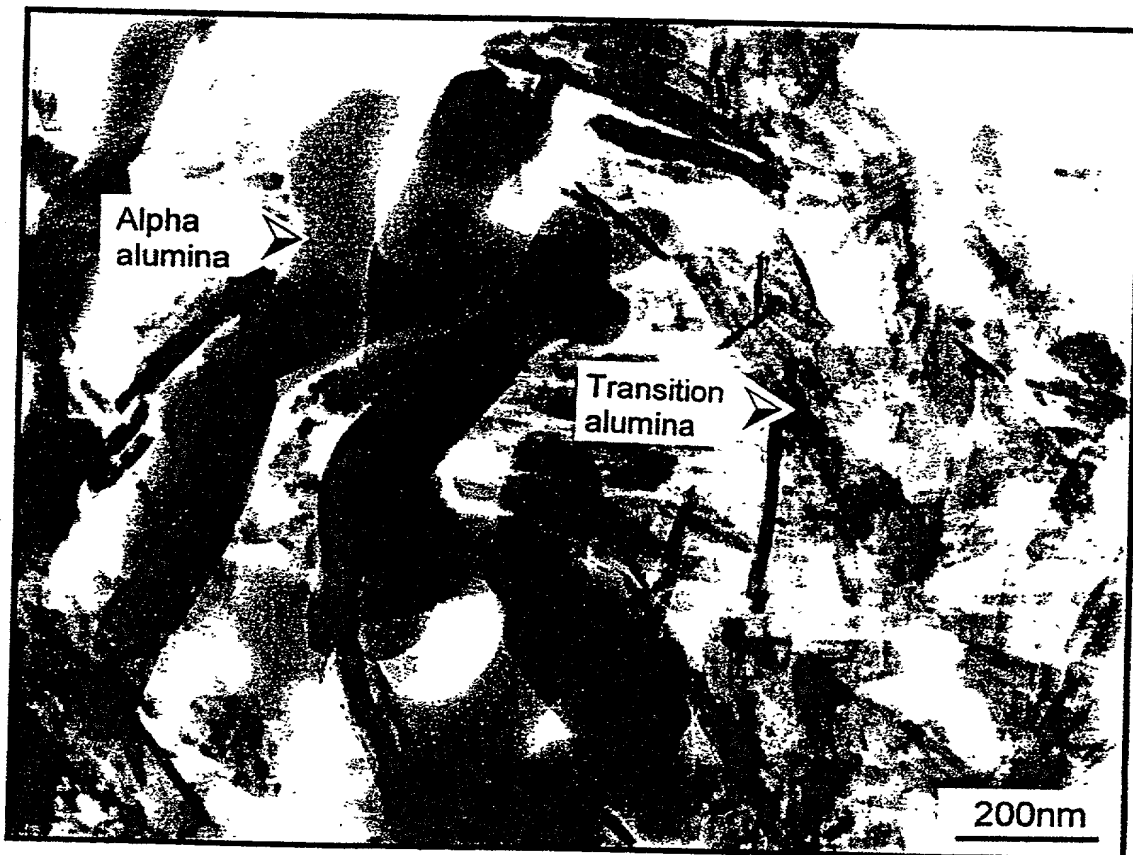


FIG. 1

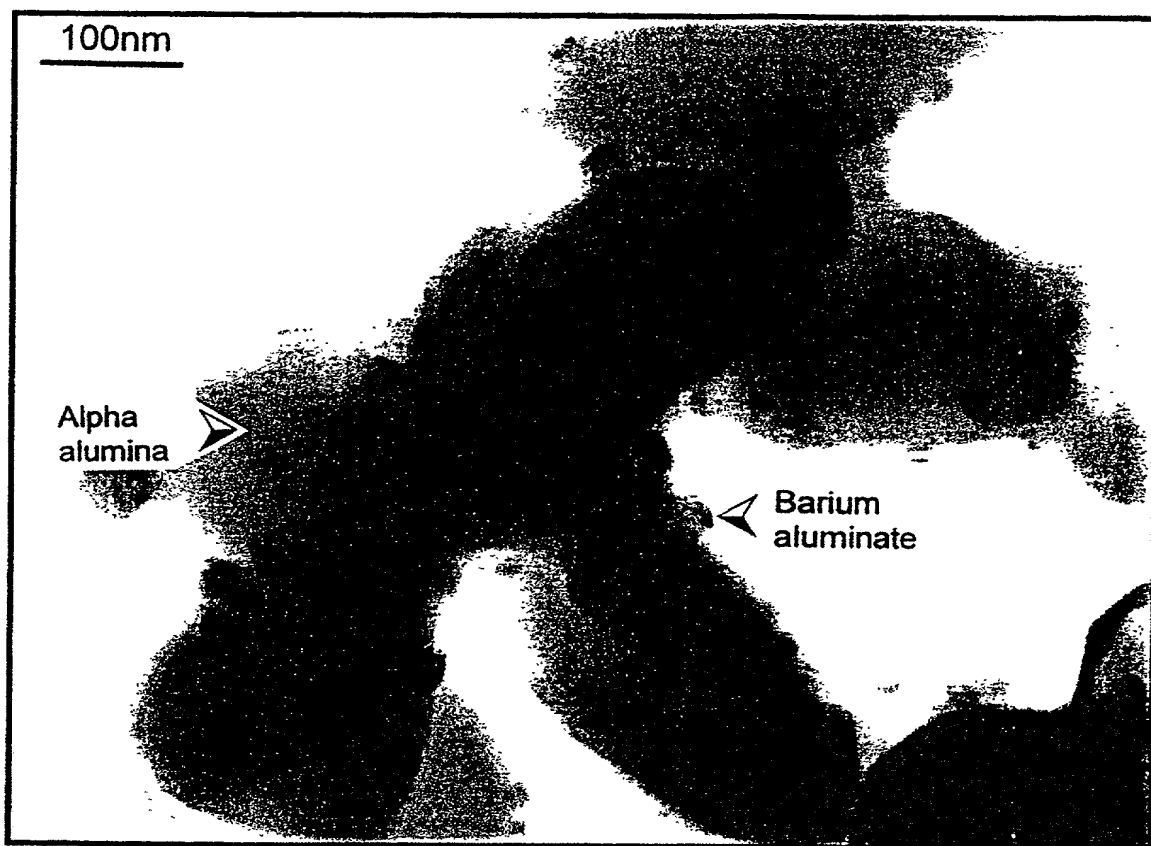


FIG. 2

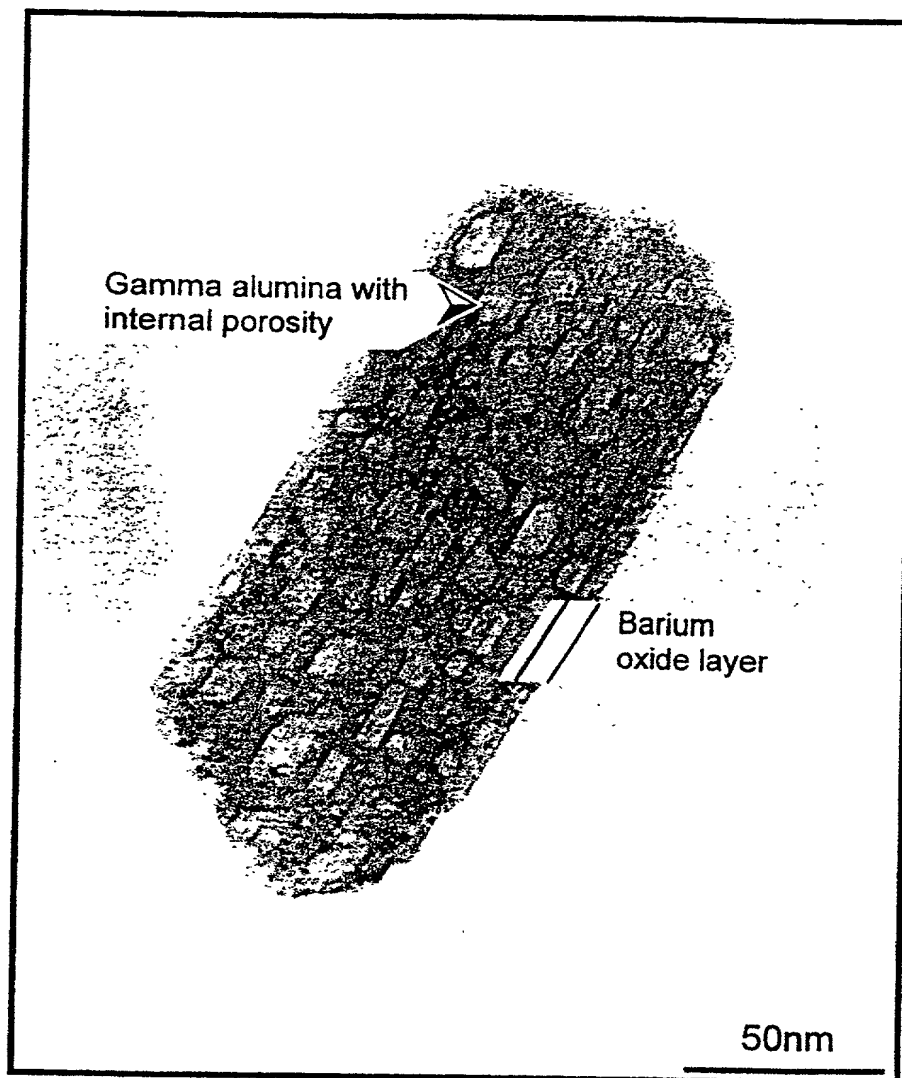


FIG. 3



FIG. 4

5 / 12



FIG. 5

6 / 12

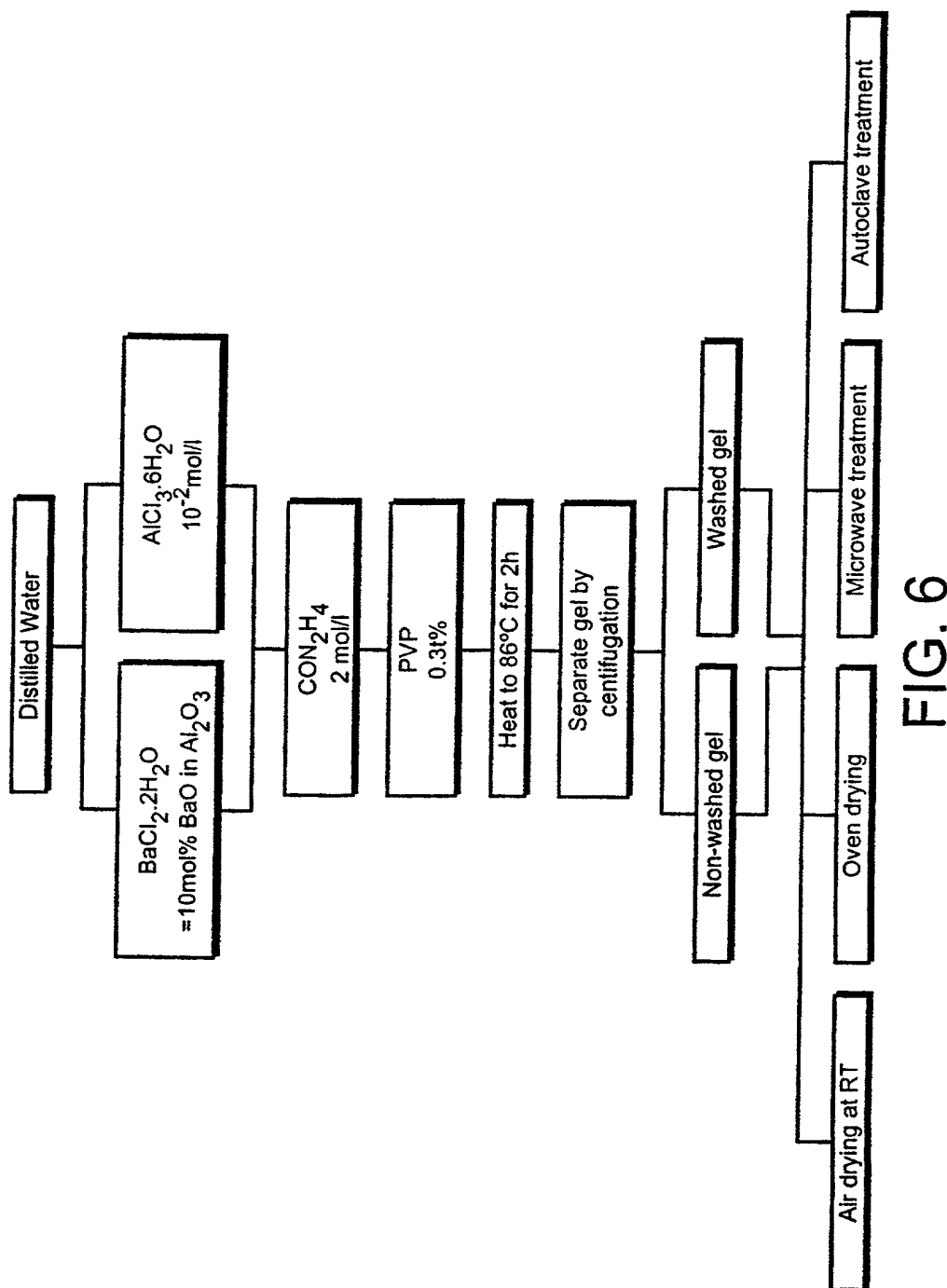


FIG. 6

7 / 12

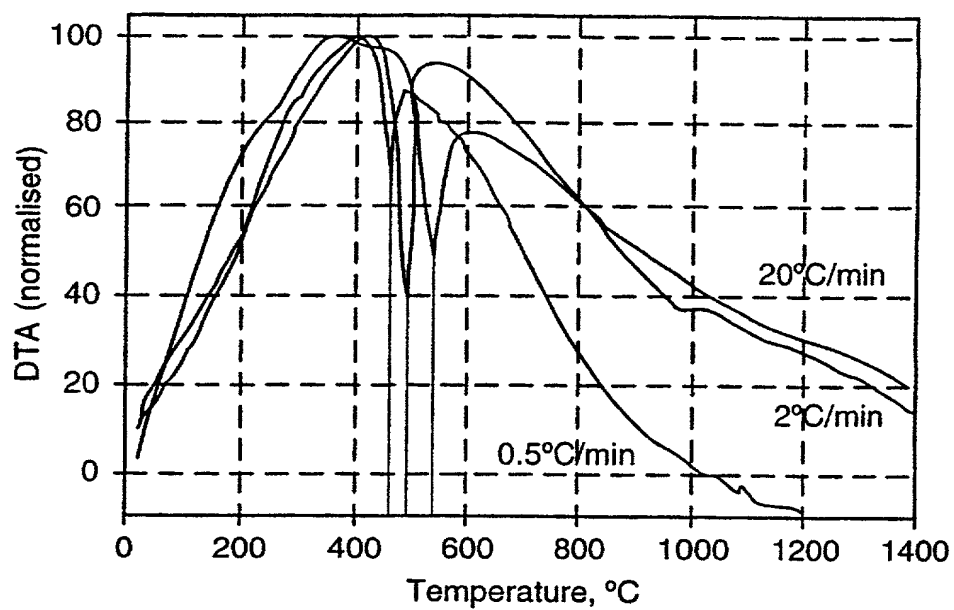


FIG. 7

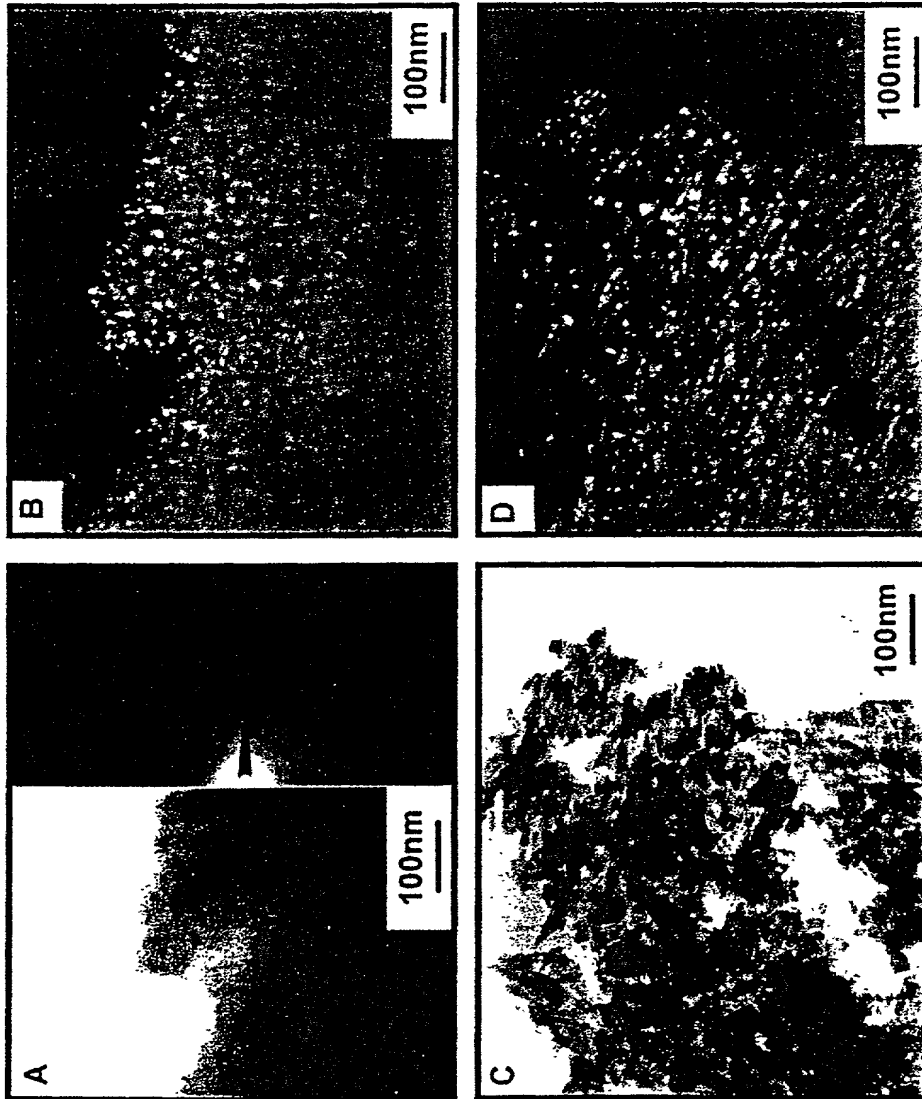


FIG. 8

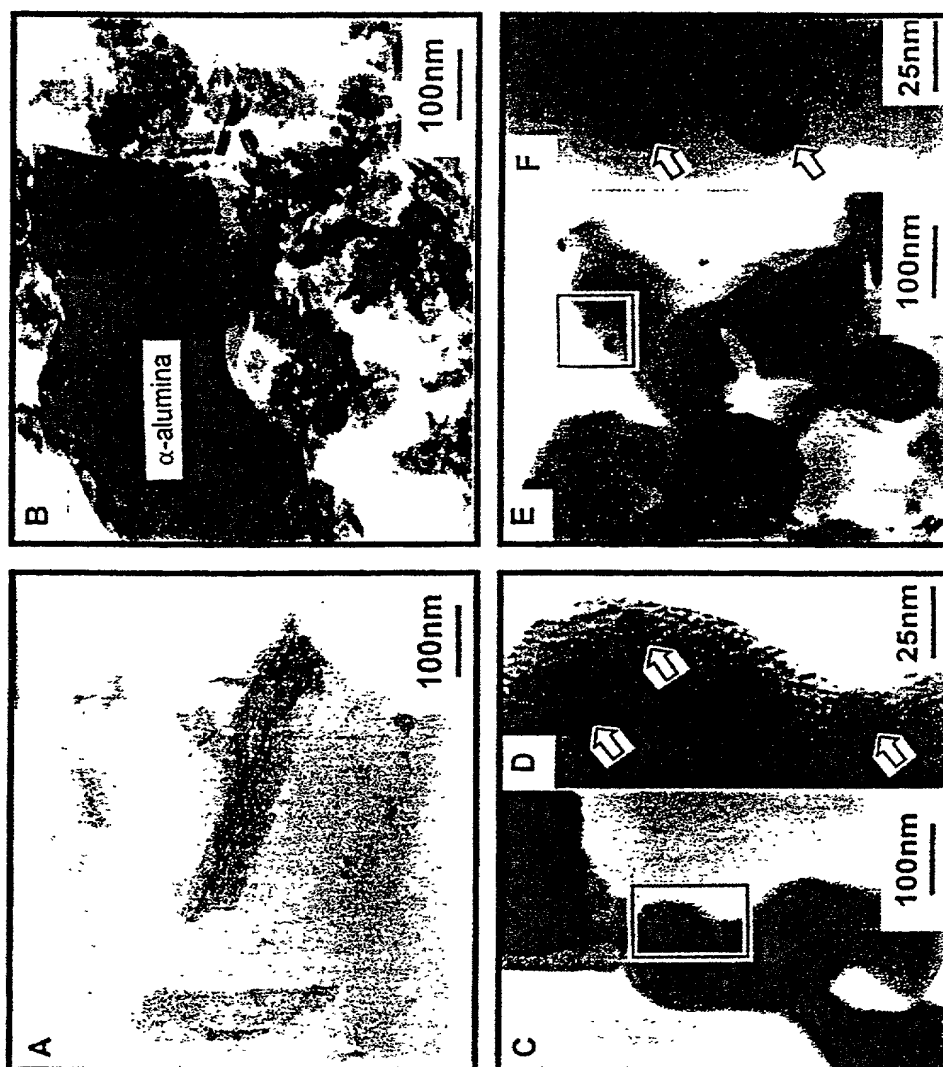


FIG. 9

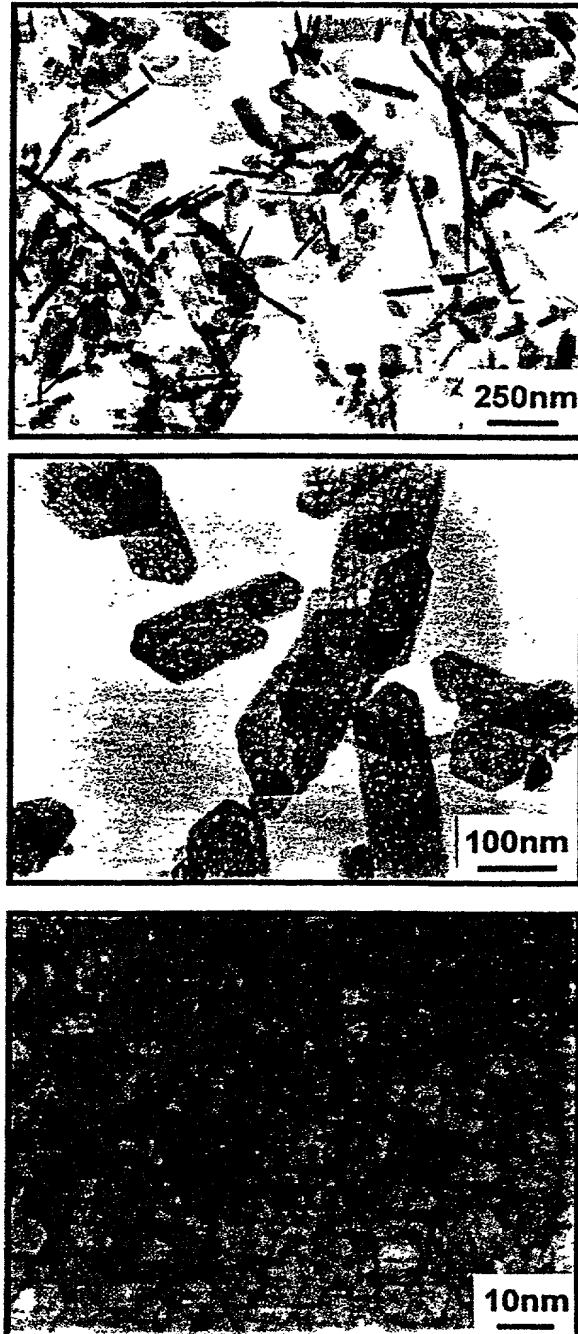


FIG. 10

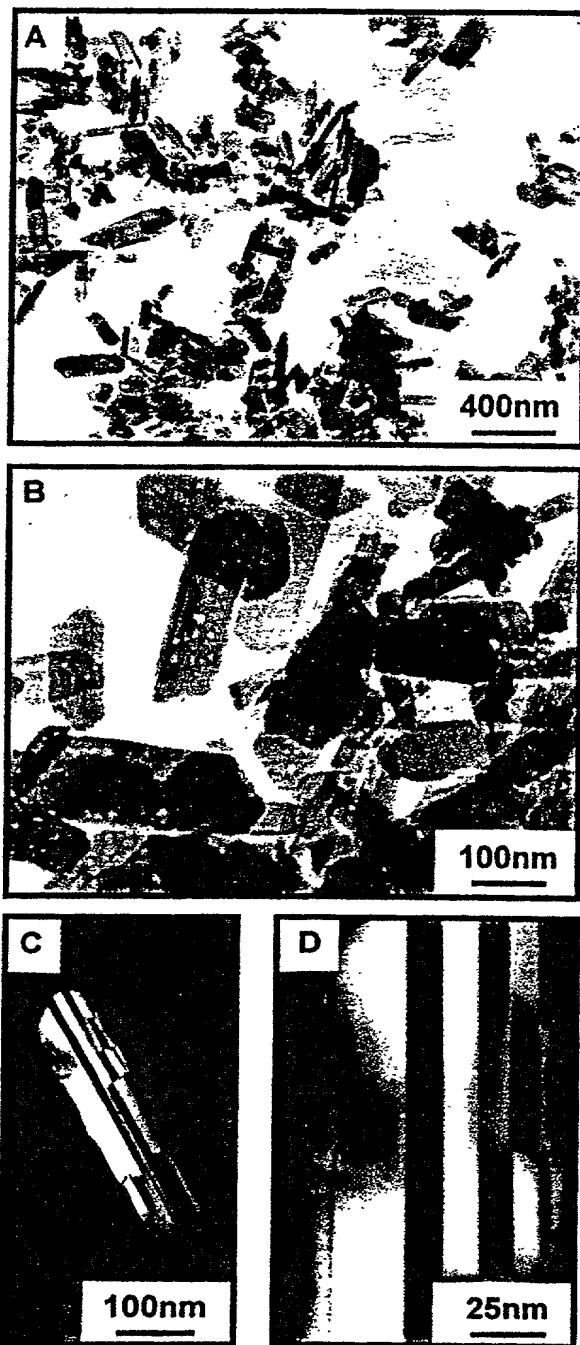


FIG. 11

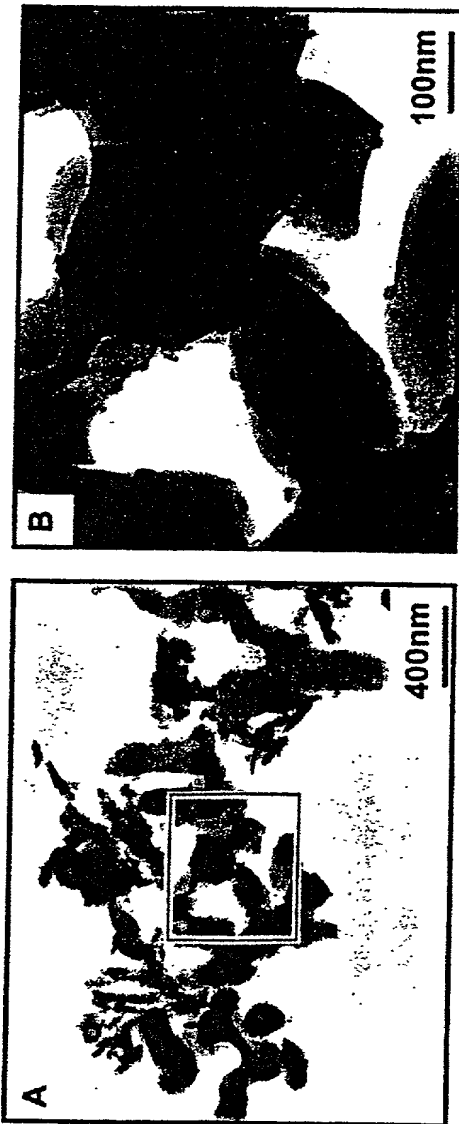


FIG. 12

DECLARATION FOR PATENT APPLICATION AND APPOINTMENT OF ATTORNEY

As a below named inventor, I hereby declare that my residence, post office address and citizenship are as stated below next to my name; I believe that I am the original, first and sole inventor (if only one name is listed below) or an original, first and joint inventor (if plural names are listed below) of the subject matter which is claimed and for which a patent is sought on the invention (Design, if applicable) entitled:

STABILIZATION OF TRANSITION ALUMINA

the specification of which (check one):

- ☐ is attached hereto, or ☒ was filed on: 17 February 2000 as U.S. Application Number or PCT International Application Number: PCT/GB00/00571 and (if applicable) was amended on:

I hereby state that I have reviewed and understand the contents of the above-identified specification, including the claims, as amended by any amendment(s) referred to above. I acknowledge the duty to disclose information which is material to patentability as defined in Title 37, Code of Federal Regulations, §1.56. I hereby claim foreign priority benefits under Title 35, United States Code §119 of any foreign application(s) for patent or inventor's certificate listed below and have also identified below any foreign application for patent or inventor's certificate having a filing date before that of the application on which priority is claimed.

PRIOR FOREIGN APPLICATION(S)			PRIORITY CLAIMED	
Number	Country	Day/Month/Year Filed	Yes	No
9904269.9	United Kingdom (GB)	24 / 2 / 1999	X	

☐ Additional Priority Application(s) Listed on Following Page(s)

I HEREBY CLAIM THE BENEFIT UNDER TITLE 35 U.S. CODE §119(E) OF ANY U.S. PROVISIONAL APPLICATIONS LISTED BELOW.	
Application Number	Day/Month/Year Filed

☐ Additional Provisional Application(s) Listed on Following Page(s)

I hereby claim the benefit under Title 35, United States Code, §120 of any United States application(s) or PCT international application(s) designating The United States of America listed below and, insofar as the subject matter of each of the claims of this application is not disclosed in that/those prior application(s) in the manner provided by the first paragraph of Title 35, United States Code, §112, I acknowledge the duty to disclose information which is material to patentability as defined in Title 37, Code of Federal Regulations, §1.56 which became available between the filing date of the prior application(s) and the national or PCT international filing date of this application:

Application Number	Filing Date	Status - Patented, Pending or Abandoned

☐ Additional US/PCT Priority Application(s) listed on Following Page(s)

I hereby declare that all statements made herein of my own knowledge are true and that all statements made on information and belief are believed to be true; and further that these statements were made with the knowledge that willful false statements and the like so made are punishable by fine or imprisonment, or both, under section 1001 of title 18 of the United States Code and that such willful false statements may jeopardize the validity of the application or any patent issued thereon.

POWER OF ATTORNEY: I (We) hereby appoint as my (our) attorneys, with full powers of substitution and revocation, to prosecute this application and transact all business in the Patent and Trademark Office connected therewith: J. Ernest Kenney, Reg. No. 19,179; Eugene Mar, Reg. No. 25,893; Richard E. Fichter, Reg. No. 26,382; Charles R. Wolfe, Jr., Reg. No. 28,680; Thomas J. Moore, Reg. No. 28,974; Bruce H. Troxell, Reg. No. 26,592; Joseph DeBenedictis, Reg. No. 28,502; Benjamin E. Urcia, Reg. No. 33,805; and

I(we) authorize my(our) attorneys to accept and follow instructions from regarding any matter related to the preparation, examination, grant and maintenance of this application, any continuation, continuation-in-part or divisional based thereon, and any patent resulting therefrom, until I(we) or my(our) assigns withdraw this authorization in writing.

Send correspondence to: **BACON & THOMAS**
625 Slaters Lane - 4th Floor
Alexandria, VA 22314

Telephone Calls to:
(703) 683-0500

FULL NAME OF FIRST OR SOLE INVENTOR Stephen Pickering	CITIZENSHIP British (GB)
RESIDENCE ADDRESS Avenue Wolfers 6 1310 La Hulpe BELGIUM	POST OFFICE ADDRESS IS THE SAME AS RESIDENCE ADDRESS UNLESS OTHERWISE SHOWN BELOW
DATE 29-10-01	SIGNATURE S. Pickering

☒ See following page(s) for additional joint inventors.

ATTORNEY/DOCKET NO:

CONTINUATION OF DECLARATION FOR PATENT APPLICATION AND APPOINTMENT OF ATTORNEY

Page 2

PRIOR FOREIGN APPLICATION(S) (35 USC §119)			PRIORITY CLAIMED	
Number	Country	Day/Month/Year Filed	Yes	No

PRIOR PROVISIONAL APPLICATIONS 35 U.S. CODE §119(e)	
Application Number	Day/Month/Year Filed

PRIOR U.S. OR PCT INTERNATIONAL APPLICATIONS (35 U.S. CODE §120)		
Application Number	Filing Date	Status - Patented, Pending or Abandoned

FULL NAME OF JOINT INVENTOR Boro Djuricic		CITIZENSHIP The Netherlands (NL)
RESIDENCE ADDRESS Myrtelstrasse 35/4/4 2340 Moedling AUSTRIA		POST OFFICE ADDRESS IS THE SAME AS RESIDENCE ADDRESS UNLESS OTHERWISE SHOWN BELOW
DATE 29 th OCTOBER 2001	SIGNATURE <i>[Signature]</i>	

FULL NAME OF JOINT INVENTOR		CITIZENSHIP
RESIDENCE ADDRESS		POST OFFICE ADDRESS IS THE SAME AS RESIDENCE ADDRESS UNLESS OTHERWISE SHOWN BELOW
DATE	SIGNATURE	

FULL NAME OF JOINT INVENTOR		CITIZENSHIP
RESIDENCE ADDRESS		POST OFFICE ADDRESS IS THE SAME AS RESIDENCE ADDRESS UNLESS OTHERWISE SHOWN BELOW
DATE	SIGNATURE	

FULL NAME OF JOINT INVENTOR		CITIZENSHIP
RESIDENCE ADDRESS		POST OFFICE ADDRESS IS THE SAME AS RESIDENCE ADDRESS UNLESS OTHERWISE SHOWN BELOW
DATE	SIGNATURE	

☒ See following pages for additional joint inventors/priority applications.

CONTINUATION OF DECLARATION FOR PATENT APPLICATION AND APPOINTMENT OF ATTORNEY

Page 3

PRIOR FOREIGN APPLICATION(S) (35 USC §119)			PRIORITY CLAIMED	
Number	Country	Day/Month/Year Filed	Yes	No

PRIOR PROVISIONAL APPLICATIONS 35 U.S. CODE §119(E)	
Application Number	Day/Month/Year Filed

PRIOR U.S. OR PCT INTERNATIONAL APPLICATIONS (35 U.S. CODE §120)		
Application Number	Filing Date	Status - Patented, Pending or Abandoned

FULL NAME OF JOINT INVENTOR	CITIZENSHIP
RESIDENCE ADDRESS	POST OFFICE ADDRESS IS THE SAME AS RESIDENCE ADDRESS UNLESS OTHERWISE SHOWN BELOW
DATE	SIGNATURE

FULL NAME OF JOINT INVENTOR <u>Jean Francois Magne</u>	CITIZENSHIP <u>FRENCH (FR)</u>
RESIDENCE ADDRESS <u>22 Rue de Puits de Denin</u> <u>16730 LINARS</u> <u>FRANCE</u>	POST OFFICE ADDRESS IS THE SAME AS RESIDENCE ADDRESS UNLESS OTHERWISE SHOWN BELOW
DATE <u>the 7th of November 2001</u>	SIGNATURE <u>Jean-Francois Magne</u>

FULL NAME OF JOINT INVENTOR	CITIZENSHIP
RESIDENCE ADDRESS	POST OFFICE ADDRESS IS THE SAME AS RESIDENCE ADDRESS UNLESS OTHERWISE SHOWN BELOW
DATE	SIGNATURE

FULL NAME OF JOINT INVENTOR	CITIZENSHIP
RESIDENCE ADDRESS	POST OFFICE ADDRESS IS THE SAME AS RESIDENCE ADDRESS UNLESS OTHERWISE SHOWN BELOW
DATE	SIGNATURE

☐ See following pages for additional joint inventors/priority applications.